

# Galvanic Cells

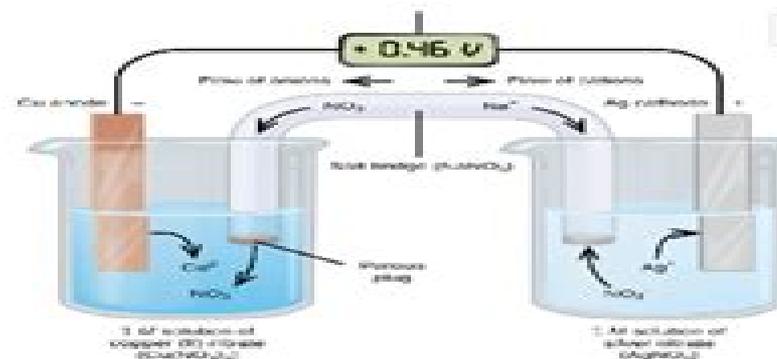
## 1. Introduction

**Aim** : The aim of this experiment is to measure the electric potential of a galvanic cell and compare it with the actual potential.

Galvanic cells are electrochemical cells in which electricity generates spontaneous oxidation reduction processes, commonly known as voltaic cells. When drafting the equations, oxidation - reduction reactions are typically separated into half-reactions to enable the entire equation to be balanced and to stress the actual chemical transitions. Oxidation is the loss of electrons or an increase in the oxidation number of an atom. Reduction is the gain of electrons or a decrease in the oxidation number of an atom.

The cell is made up of two half cells. The half cell shown at the left of the figure contains copper in an aqueous solution of copper nitrate. The half cell on the right contains silver ions in an aqueous solution of silver nitrate. A tube containing an inert electrolyte solution, known as a **salt bridge** is used to separate the two half cells. The electrodes in an electrochemical cell are called the cathode and anode. They are made of a solid metal and are submerged into a solution of the electrodes' ion called the electrolyte. For example, in Figure 1 since the electrode is made up of copper, the electrolyte is in a solution of an  $\text{Cu}^{2+}$  ions. **The anode** is where the oxidation occurs. It is always drawn on the left. **The cathode** is where the reduction takes place. It is always drawn on the right. An external conductor is connected to both the electrodes which completes the circuit.

FIGURE 1. Schematics of a galvanic cell.



As oxidation takes place in the anode, Zn oxidises and turns into  $\text{Zn}^{2+}$

# Galvanic Cell Lab Report

**N Noddings**



## **Galvanic Cell Lab Report:**

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## **Table of Contents Galvanic Cell Lab Report**

1. Understanding the eBook Galvanic Cell Lab Report
  - The Rise of Digital Reading Galvanic Cell Lab Report
  - Advantages of eBooks Over Traditional Books
2. Identifying Galvanic Cell Lab Report
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Galvanic Cell Lab Report
  - User-Friendly Interface
4. Exploring eBook Recommendations from Galvanic Cell Lab Report
  - Personalized Recommendations
  - Galvanic Cell Lab Report User Reviews and Ratings
  - Galvanic Cell Lab Report and Bestseller Lists
5. Accessing Galvanic Cell Lab Report Free and Paid eBooks
  - Galvanic Cell Lab Report Public Domain eBooks
  - Galvanic Cell Lab Report eBook Subscription Services
  - Galvanic Cell Lab Report Budget-Friendly Options
6. Navigating Galvanic Cell Lab Report eBook Formats

- ePub, PDF, MOBI, and More
  - Galvanic Cell Lab Report Compatibility with Devices
  - Galvanic Cell Lab Report Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Galvanic Cell Lab Report
    - Highlighting and Note-Taking Galvanic Cell Lab Report
    - Interactive Elements Galvanic Cell Lab Report
  8. Staying Engaged with Galvanic Cell Lab Report
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Galvanic Cell Lab Report
  9. Balancing eBooks and Physical Books Galvanic Cell Lab Report
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Galvanic Cell Lab Report
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Galvanic Cell Lab Report
    - Setting Reading Goals Galvanic Cell Lab Report
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Galvanic Cell Lab Report
    - Fact-Checking eBook Content of Galvanic Cell Lab Report
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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