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web balancing equations and simple stoichiometry answers balance the following equations  $1\text{ N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$   $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$   $3\text{HBr} + 1\text{KHCO}_3 \rightarrow 1\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2$   $4\text{Ga} + 3\text{Na}_2\text{SO}_3 \rightarrow 2\text{Ga}_2\text{SO}_3 + 3\text{S} + 6\text{NaBr}$   $5\text{S} + 3\text{NO}_2 \rightarrow 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$

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web google classroom balance the following chemical equation  $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$   $\text{HCl} \rightarrow \text{HCl}$   $\text{MgCl}_2$   $\text{H}_2$   $\text{O}$   $\text{H}_2\text{O}$  note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem  $7\text{H}_4\text{X} + \text{Y} \rightarrow \theta\text{PI} + \pi\text{H}_8 + 5\text{H}_2 + 0\text{H}_9 + 6$

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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