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Question 3 (continued)

(c) The Montreal Protocol of 1987 provided a global framework to phase out chlorofluorocarbon (CFC) production and use. Although the Montreal Protocol has led to a dramatic decrease in CFCs released into the atmosphere, stratospheric ozone destruction has decreased only slightly.

(i) Explain the process by which CFCs lead to the destruction of stratospheric ozone. (You may use chemical equations in your answer.)

*(2 points: 1 point for a correct description of the decomposition of CFCs and 1 point for a correct description of the reaction of ozone with chlorine)*

Correct descriptions of the decomposition of CFCs include one of the following:

- Absorption of UV radiation by CFC molecules releases chlorine atoms
- $\text{CCl}_2\text{F} + \text{UV} \rightarrow \text{CCl}_2\text{F} + \text{Cl}$

Correct descriptions of the destruction of stratospheric ozone include one of the following:

- Chlorine atoms break down ozone molecules
- $\text{Cl} + \text{O}_3 \rightarrow \text{ClO} + \text{O}_2$

(ii) Explain why the rapid decrease in CFC emissions has not led to a similarly rapid decrease in the destruction of stratospheric ozone.

*(1 point can be earned for a correct explanation linking the absence of a rapid decrease in the destruction of stratospheric ozone with one of the following):*

- The slow migration of CFCs into the stratosphere.
- The long lifetime of CFCs and/or chlorine in the stratosphere.
- The continued release of other ozone-depleting substances.

(d) Identify a human activity that leads to the formation of tropospheric ozone as a secondary pollutant and explain why tropospheric ozone levels peak in the daytime.

*(2 points: 1 point for a correct human activity and 1 point for correctly explaining that sunlight is required to form tropospheric ozone; only the first human activity can earn points)*

The following are acceptable human activities:

- Burning fossil fuels (e.g., operating motor vehicles, using landscaping equipment, electric power generation, industrial production)
- Releasing VOCs (e.g., pumping gas, using solvent-based paints)

The following are acceptable explanations for why tropospheric ozone levels peak in the daytime:

- Sunlight is required to form tropospheric ozone
- Tropospheric ozone is created by photochemical reactions

# Environmental Science 2013 Response Scoring Guidelines

**A Gutmann**



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