

## General Chemistry I Laboratory Report

Name: \_\_\_\_\_

Lab day: \_\_\_\_\_

Date of experiment: \_\_\_\_\_

TA: \_\_\_\_\_

**10. Redox reactions and Electrochemical cells**Purposes of the experiment and brief overview of experimental method(s)Results and Observations:Part I qualitative observations

For the tests you carried out, write the observations you made, and write balanced net ionic reactions below. If nothing happened in a test, write 'no reaction'.

Reactants	Observations
$Zn + Cu^{2+}$	Zinc turns black. $Zn + Cu^{2+} \rightarrow Cu + Zn^{2+}$
$Zn^{2+} + Cu$	No reaction. $Cu + Zn^{2+} \rightarrow Zn + Cu^{2+}$
$Zn^{2+} + Pb$	No reaction. $Pb + Zn^{2+} \rightarrow Zn + Pb^{2+}$
$Zn + Pb^{2+}$	Zinc turns black. $Zn + Pb^{2+} \rightarrow Pb + Zn^{2+}$
$Pb^{2+} + Cu$	No reaction. $Cu + Pb^{2+} \rightarrow Pb + Cu^{2+}$
$Pb + Cu^{2+}$	Lead metal turns black. $Pb + Cu^{2+} \rightarrow Zn + Cu^{2+}$
$Zn + HCl$	Zinc reacts with hydrochloric acid to form bubbles.

# General Chemistry Lab Manual Answers Redox Lab

**Victor M. Corman**



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