

CHAPTER 14: COVALENT BONDING: ORBITALS

PROBLEMS WITH LEWIS STRUCTURE

- Provides no indication of the **shapes** of the molecules
- O_2 would be predicted to be diamagnetic, but it is **paramagnetic**

Paramagnetic = has unpaired electrons.



The Lewis structure of O_2 has an $O=O$ double bond with no unpaired electrons, i.e. diamagnetic.

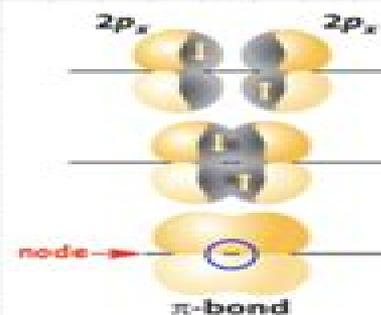
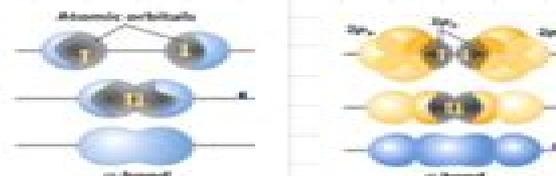
- Many compounds cannot be represented by a single structure
- Resonance structures **differ only in position of electrons** not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
- Can't explain molecules such as B_2H_6 (electron deficient molecules)

VALENCE-BOND THEORY "LOCALIZED ELECTRON MODEL"

- Covalent bond is created by the overlap of orbitals on two atoms

TYPES OF OVERLAP: σ and π Bonds

- **Bonding in H_2**
 - Sausage shaped orbital, bond called σ -orbital
 - Called σ orbital because it has **no node** and is **cylindrically symmetrical**
 - Electron density is **concentrated between two nuclei**, but is still highest at the nuclei themselves.
- **Bonding in N_2**
 - σ orbital is obtained by overlap of **two p_z -orbitals**
- Perpendicular p orbitals overlap in a side-by-side manner.
- The resulting orbital has two lobes of electron density and is called a **π -bond**.
- Called π -orbital because it **has a node** and, when viewed down the internuclear axis, it looks like a p-orbital
- π -bond has **no electron density between the nuclei along the bond axis**.



General Chemistry Chapter 14 Covalent Bonding Orbitals Questions

J Ma



General Chemistry Chapter 14 Covalent Bonding Orbitals Questions:

Reviewing **General Chemistry Chapter 14 Covalent Bonding Orbitals Questions**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**General Chemistry Chapter 14 Covalent Bonding Orbitals Questions**," an enthralling opus penned by a highly acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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