

General Chemistry II Jasperse  
Kinetics. Extra Practice Problems

General Types/Groups of problems:

Rates of Change in Chemical Reactions	p1	First Order Rate Law Calculations	P9
The look of concentration/time graphs	p2	Reaction Energy Diagrams, Activation Energy, Transition States...	P10
Rates: Average Rates, Determination of Rates from Stoichiometry and Changes of Other Chemicals	p3	Reaction Mechanisms, Intermediates...	P12
Reactant Order and Overall Reaction Order	P4	Mechanism Steps and Rate Laws	P13
Given a Rate Law, How much will rate change with change in concentration	P5	Catalysts	P14
Determining Reactant Orders from Actual Data	P6	Answers	P15
Actual Rate Law from Concentration/Rate Data	P7		
Determining Rate Constant from Actual Data. Where you yourself need to first deduce the rate law, then plug in the values to solve for "k".	P8		

**Rates of Change in Chemical Reactions**

- For the reaction  $A + 3B \rightarrow 2C$ , how does the rate of disappearance of B compare to the rate of production of C?
  - the rate of disappearance of B is 1/2 the rate of appearance of C
  - the rate of disappearance of B is 3/2 the rate of appearance of C
  - the rate of disappearance of B is 2/3 the rate of appearance of C
  - the rate of disappearance of B is 1/3 the rate of appearance of C
- For the reaction  $2A + 3B \rightarrow 4C + 5D$ , the rate of the reaction in terms of  $\Delta A$  would be written as:
  - $-\Delta A/\Delta t$
  - $-1/2 \Delta A/\Delta t$
  - $+\Delta A/\Delta t$
  - $+1/2 \Delta A/\Delta t$
  - $-2 \Delta A/\Delta t$
- For the reaction  $2A + 3B \rightarrow 4C + 5D$ , the rate of the reaction in terms of  $\Delta B$  would be written as:
  - $-\Delta B/\Delta t$
  - $+\Delta B/\Delta t$
  - $-1/3 \Delta B/\Delta t$
  - $+1/3 \Delta B/\Delta t$
  - $-3 \Delta B/\Delta t$
- For the reaction  $2A + 3B \rightarrow 4C + 5D$ , the rate of the reaction in terms of  $\Delta C$  would be written as:
  - $+\Delta C/\Delta t$
  - $+4 \Delta C/\Delta t$
  - $+1/4 \Delta C/\Delta t$
  - $-4 \Delta C/\Delta t$
  - $-1/4 \Delta C/\Delta t$
- In the combustion of methane,  $CH_4(g) + 2 O_2(g) \rightarrow CO_2(g) + 2 H_2O(g)$ , which reactant has the greatest rate of disappearance?
  - $CH_4$
  - $O_2$
  - $CO_2$
  - $H_2O$
  - $CH_4$  and  $O_2$  have the same rate of disappearance.

# General Chemistry II Chapter Chemical Kinetics Practice Problems

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### **Table of Contents General Chemistry Ii Chapter Chemical Kinetics Practice Problems**

1. Understanding the eBook General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - The Rise of Digital Reading General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Personalized Recommendations
  - General Chemistry Ii Chapter Chemical Kinetics Practice Problems User Reviews and Ratings
  - General Chemistry Ii Chapter Chemical Kinetics Practice Problems and Bestseller Lists

5. Accessing General Chemistry Ii Chapter Chemical Kinetics Practice Problems Free and Paid eBooks
  - General Chemistry Ii Chapter Chemical Kinetics Practice Problems Public Domain eBooks
  - General Chemistry Ii Chapter Chemical Kinetics Practice Problems eBook Subscription Services
  - General Chemistry Ii Chapter Chemical Kinetics Practice Problems Budget-Friendly Options
6. Navigating General Chemistry Ii Chapter Chemical Kinetics Practice Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - General Chemistry Ii Chapter Chemical Kinetics Practice Problems Compatibility with Devices
  - General Chemistry Ii Chapter Chemical Kinetics Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Highlighting and Note-Taking General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Interactive Elements General Chemistry Ii Chapter Chemical Kinetics Practice Problems
8. Staying Engaged with General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers General Chemistry Ii Chapter Chemical Kinetics Practice Problems
9. Balancing eBooks and Physical Books General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection General Chemistry Ii Chapter Chemical Kinetics Practice Problems
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Setting Reading Goals General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Fact-Checking eBook Content of General Chemistry Ii Chapter Chemical Kinetics Practice Problems
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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