

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- | | |
|-----------------------------|---|
| a. Sulfur hexafluoride | SF_6 |
| b. Lithium nitride | Li_3N |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$ |
| d. Tin (II) chloride | SnCl_2 |
| e. Ammonium acetate | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$) |
| f. Mercury (I) chloride | Hg_2Cl_2 |
| g. Potassium bromate | KBrO_3 |
| h. Hydrosulfuric acid | $\text{H}_2\text{S(aq)}$ |
| i. Chloric acid | HClO_3 |
| j. Sulfurous acid | H_2SO_3 |

Name each of the following compounds:

- | | |
|------------------------------|-------------------------|
| a. CuSO_4 | Copper (II) sulfate |
| b. AlF_3 | aluminum fluoride |
| c. HI | hydrogen iodide |
| d. NO | nitrogen monoxide |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid |
| f. HNO_3 | nitric acid |
| g. HNO_2 | nitrous acid |
| h. NaHSO_4 | Sodium hydrogen sulfate |

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 44 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ + \\ \hline 63.62 \rightarrow 64 \end{array}$$

Health Chemistry Chapter Review Answers

Emilie Sanchez



Health Chemistry Chapter Review Answers:

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Table of Contents Health Chemistry Chapter Review Answers

1. Understanding the eBook Health Chemistry Chapter Review Answers
 - The Rise of Digital Reading Health Chemistry Chapter Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Health Chemistry Chapter Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Health Chemistry Chapter Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Health Chemistry Chapter Review Answers

- Personalized Recommendations
 - Health Chemistry Chapter Review Answers User Reviews and Ratings
 - Health Chemistry Chapter Review Answers and Bestseller Lists
5. Accessing Health Chemistry Chapter Review Answers Free and Paid eBooks
 - Health Chemistry Chapter Review Answers Public Domain eBooks
 - Health Chemistry Chapter Review Answers eBook Subscription Services
 - Health Chemistry Chapter Review Answers Budget-Friendly Options
 6. Navigating Health Chemistry Chapter Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Health Chemistry Chapter Review Answers Compatibility with Devices
 - Health Chemistry Chapter Review Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Health Chemistry Chapter Review Answers
 - Highlighting and Note-Taking Health Chemistry Chapter Review Answers
 - Interactive Elements Health Chemistry Chapter Review Answers
 8. Staying Engaged with Health Chemistry Chapter Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Health Chemistry Chapter Review Answers
 9. Balancing eBooks and Physical Books Health Chemistry Chapter Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Health Chemistry Chapter Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Health Chemistry Chapter Review Answers
 - Setting Reading Goals Health Chemistry Chapter Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Health Chemistry Chapter Review Answers

- Fact-Checking eBook Content of Health Chemistry Chapter Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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