

Fill in the table on the right from your periodic table.

Element	Chemical Symbol	Atomic #	Valence Electrons	Oxidation #	# of Protons	Atomic Mass	Metal or Nonmetal	Group or Luster Electrons	# of Full Electron Levels
Oxygen									
Argon									
Calcium									
Nitrogen									
Potassium									
Sulfur									
Helium									

Find the molar mass of the following compounds

Calcium oxide (CaO)

Water

Classify (ionic, covalent or polyatomic) and name the following compounds.

- CO₂ _____
- N₂O _____
- NaClO₂ _____

Draw the Lewis Dot Diagrams for

- Carbon
- Sodium

Make balanced ionic compounds from:

- Li + O _____
- Be + NO₃⁻ _____
- Na + Ar _____
- Ca + O _____

With dot diagrams draw the covalent bond of O₂

With dot diagrams the covalent bond of CH₄

Matching:

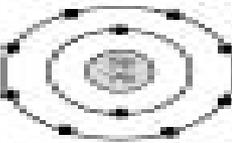
- Mixture
- Compound
- Element
- Matter
- Heterogeneous
- Homogeneous

- _____ A substance made up of two or more atoms that must be separated by chemical means.
- _____ Something made up of 2 or more different atoms that can be separated by physical means.
- _____ A substance that is pure and made up of only one type of atom.
- _____ Anything that has mass and takes up space.
- _____ A mixture that is the same throughout.
- _____ A mixture that is different throughout.

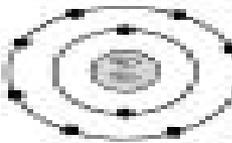
1. Valence electrons are the outermost electrons of an atom that are involved in chemical bonding. True/False

2. Lithium has _____ valence electrons, will _____ (lose or gain) _____ electrons and become _____ (positive or negative). This is why lithium's oxidation number is _____.

3. A metal and non-metal will form a _____ compound, two non-metals form a _____ compound.



Atom A



Atom B

Which one is the ion?

Give the ion notation for the ion.

Matching:

- Ion
- Element
- Isotope
- Oil

- _____ If you change the number of protons, you change the _____.
- _____ If you change the number of neutrons, you change the _____.
- _____ If you change the number of electrons, you change the _____.
- _____ Every 3,000 miles you should change a car's _____.

Matching:

- Unit of Measurement of Mass
- Meter
- Conductor
- Leaf
- Claw

- _____ Unit of length; equal to 3.3 ft.
- _____ Mass is neither destroyed or created in chemical reactions.
- _____ Unit of area; about 1 dollar bill.
- _____ 1/100th of mass; width of pinky finger.
- _____ Unit of volume; just bigger than a quart.

Matching:

- Solid
- Liquid
- Gas
- Oil

- _____ Molecules that are tightly packed and retain their shape and size.
- _____ Molecules that bounce off of each other, can be compressed, and take the shape of their container.
- _____ Molecules that can move (slide) over each other and have a definite size (volume), but not shape and can not be compressed.

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Table of Contents Fall Semester Review Ipc Answers 2014

1. Understanding the eBook Fall Semester Review Ipc Answers 2014
 - The Rise of Digital Reading Fall Semester Review Ipc Answers 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Fall Semester Review Ipc Answers 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Fall Semester Review Ipc Answers 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Fall Semester Review Ipc Answers 2014

- Personalized Recommendations
 - Fall Semester Review Ipc Answers 2014 User Reviews and Ratings
 - Fall Semester Review Ipc Answers 2014 and Bestseller Lists
5. Accessing Fall Semester Review Ipc Answers 2014 Free and Paid eBooks
- Fall Semester Review Ipc Answers 2014 Public Domain eBooks
 - Fall Semester Review Ipc Answers 2014 eBook Subscription Services
 - Fall Semester Review Ipc Answers 2014 Budget-Friendly Options
6. Navigating Fall Semester Review Ipc Answers 2014 eBook Formats
- ePub, PDF, MOBI, and More
 - Fall Semester Review Ipc Answers 2014 Compatibility with Devices
 - Fall Semester Review Ipc Answers 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Fall Semester Review Ipc Answers 2014
 - Highlighting and Note-Taking Fall Semester Review Ipc Answers 2014
 - Interactive Elements Fall Semester Review Ipc Answers 2014
8. Staying Engaged with Fall Semester Review Ipc Answers 2014
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Fall Semester Review Ipc Answers 2014
9. Balancing eBooks and Physical Books Fall Semester Review Ipc Answers 2014
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Fall Semester Review Ipc Answers 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Fall Semester Review Ipc Answers 2014
- Setting Reading Goals Fall Semester Review Ipc Answers 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Fall Semester Review Ipc Answers 2014

- Fact-Checking eBook Content of Fall Semester Review Ipc Answers 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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