

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, LeChatelier's principle and equilibrium constant applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{Fe}^{3+}(\text{aq}) + \text{HSCN}(\text{aq})$$
4. In part 4, what is the conjugate acid of the base SCN^{-} and what is the conjugate base of the acid HSCN ?
Conjugate acid: HSCN (the central group and hence the proton donor) and conjugate base: SCN^{-} (the proton acceptor) (the central group and hence the proton acceptor).
5. What is the conjugate acid of the base SCN^{-} in part 4?
Conjugate acid: HSCN (the central group and hence the proton donor).
6. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$
8. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$
9. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$

Investigating Chemical Equilibrium Lab 12a

LL Leslie



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