

## GRAPHING PERIODIC TRENDS

**Standard:** Students know how to use the periodic table to identify trends in ionization energy, electronegativity, and the relative sizes of ions and atoms.

### PRE-LAB DISCUSSION:

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern. Students can discover these patterns by examining the changes in properties of elements on the Periodic Table. The properties that will be examined in this lesson are: atomic radius AND first ionization energy.

### PURPOSE:

To understand periodic trends relating to atomic radius, ionization energy, and electronegativity.

### PROCEDURE:

Use the information in these tables to complete the graph as described below.

### RESULTS:

| Symbol | Atomic Radius (Picometers) | First Ionization (Kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|------------------------------------|-----------------------------------|
| H      | 37                         | 1312                               | 2.1                               |
| Li     | 128                        | 520                                | 1.0                               |
| Na     | 190                        | 495                                | 0.9                               |
| K      | 203                        | 419                                | 0.8                               |
| Rb     | 220                        | 403                                | 0.8                               |
| Cs     | 244                        | 376                                | 0.7                               |

| Symbol | Atomic Radius (Picometers) | First Ionization (Kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|------------------------------------|-----------------------------------|
| Na     | 190                        | 495                                | 0.9                               |
| Mg     | 141                        | 738                                | 1.2                               |
| Al     | 121                        | 578                                | 1.5                               |
| Si     | 111                        | 787                                | 1.8                               |
| P      | 107                        | 1012                               | 2.1                               |
| S      | 104                        | 1000                               | 2.5                               |
| Cl     | 102                        | 1251                               | 3.0                               |
| Ar     | 100                        | 1521                               | —                                 |

1. Define each of these terms: Atomic radius, First Ionization Energy, and Electronegativity
2. Using colored pencil or pen, list under the symbol (in this order) the Atomic Radius, First Ionization Energy, and Electronegativity. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First Ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)
5. There is NO ABSTRACT to write for this activity :)

# Graphing Periodic Relationships Lab Answers

**DP Hallahan**



## **Graphing Periodic Relationships Lab Answers:**

## Unveiling the Power of Verbal Art: An Mental Sojourn through **Graphing Periodic Relationships Lab Answers**

In a global inundated with displays and the cacophony of immediate interaction, the profound power and emotional resonance of verbal beauty often diminish into obscurity, eclipsed by the constant assault of sound and distractions. However, situated within the musical pages of **Graphing Periodic Relationships Lab Answers**, a charming perform of literary beauty that impulses with raw feelings, lies an unique trip waiting to be embarked upon. Written with a virtuoso wordsmith, this mesmerizing opus instructions visitors on a psychological odyssey, gently revealing the latent potential and profound impact embedded within the intricate web of language. Within the heart-wrenching expanse of this evocative examination, we can embark upon an introspective exploration of the book is main themes, dissect their captivating publishing fashion, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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