

CHAPTER  
10

## Review

## More Equations and Inequalities

## 10-1 Solving Two-Step Equations (pp. 498–501)

Solve.

1.  $2a + 8 = 16$

$$a = 4$$

4.  $\frac{4+t}{6} = 2$

$$t = 8$$

2.  $6 - 5p = -24$

$$p = 6$$

5.  $\frac{m}{2} - 3 = 3$

$$m = 12$$

3.  $-4x - 2 = 2$

$$x = -1$$

6.  $10 = \frac{x+1}{2}$

$$x = 19$$

7. The bill to repair a car was \$439. The cost for parts was \$265. The cost for the labor was \$29 per hour. How many hours did the repair work take?

$$6 \text{ hours}$$

## 10-2 Solving Multi-Step Equations (pp. 502–505)

Solve.

8.  $-4v - 20 + 2v = -40$

$$v = 10$$

9.  $\frac{3x}{2} + \frac{2x}{4} = 8$

$$x = 4$$

10.  $\frac{c}{10} - \frac{6}{10} = \frac{3}{5}$

$$c = 12$$

11.  $\frac{4k}{3} + \frac{k}{6} = \frac{5}{4}$

$$k = \frac{5}{6} \approx 0.83$$

12.  $4 - 4s - 6 + 14s = 18$

$$s = 2$$

13.  $\frac{5}{2}n + \frac{2}{4}n = 3$

$$n = 1$$

- ~~14. The average running time of the top 4 finishers in a road race was 18.80 min. The times of the second-, third-, and fourth-place finishers were 18.45 min, 19.27 min, and 19.80 min. What was the time of the first-place finisher?~~

## 10-3 Solving Equations with Variables on Both Sides (pp. 507–512)

Solve.

15.  $4s + 3 = 5s + 4$

$$s = -1$$

16.  $12h - 5 = 5(2h + 1)$

$$h = 5$$

17.  $\frac{5d}{4} - \frac{15}{3} = 2d - \frac{5}{4}$

$$d = -5$$

18. Vertical angles are congruent. Find the measure of the angles.

$$67^\circ$$



# Math Madness Answer Key

**JA Banks**



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web 1 answer a explanation the value of  $\Delta h$  is negative therefore the reaction is exothermic or heat releasing 2 answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta h$  must also be halved 3 answer c explanation option c is incorrect because 566 kj mol is the  $\Delta h$  per mole of  $\text{o}_2$  and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes  $\text{h}_2\text{o (s)} \rightarrow \text{h}_2\text{o (l)}$   $6.01 \text{ kJ/mol}$  this is an endothermic process it requires 6.01 kJ to melt one mole of ice  $\text{h}_2\text{o (s)}$  the enthalpy value will change if the number of moles varies from the 1:1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard

enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy  $e$  of a system in any process is equal to the heat  $q$  added to the system plus the work  $w$  done on the system by its surroundings  $e$

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