

Answer KeyNaming and Writing Formulas for Ionic Compounds

Write the names for the following compounds:

1. KClO_4 – potassium perchlorate
2. Sr_3P_2 – strontium phosphide
3. NH_4NO_2 – ammonium nitrite
4. NaCl – sodium chloride
5. CoS – cobalt (II) sulfide
6. ZnCl_2 – zinc (II) chloride
7. Li_3N – lithium nitride
8. BaSO_3 – barium sulfite
9. Co_2S_3 – cobalt (III) sulfide
10. CsF – cesium fluoride
11. NaNO_3 – sodium nitrate
12. CuI_2 – copper (II) iodide
13. TiO_2 – titanium (IV) oxide
14. KMnO_4 – potassium permanganate
15. ZnCN_2 – zinc (II) cyanide
16. $\text{Al}_2(\text{CO}_3)_3$ – aluminum carbonate
17. $\text{Fe}_2(\text{CO}_3)_3$ – iron (III) carbonate
18. BaCrO_4 – barium chromate
19. $\text{Pb}_3(\text{PO}_4)_2$ – lead (II) phosphate
20. NaHCO_3 – sodium hydrogen carbonate or sodium bicarbonate

Write the formulas for the following compounds:

1. cesium bromide – CsBr
2. aluminum chlorite – $\text{Al}(\text{ClO}_2)_3$
3. ammonium phosphite – $(\text{NH}_4)_3\text{PO}_3$
4. copper (II) nitrate – $\text{Cu}(\text{NO}_3)_2$
5. sodium fluoride – NaF
6. lead (IV) sulfate – $\text{Pb}(\text{SO}_4)_2$
7. calcium carbonate – CaCO_3
8. barium oxide – BaO
9. iron (III) sulfide – Fe_2S_3
10. manganese (II) chromate – MnCrO_4

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J Ma



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Science Spectrum Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03 **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Carolina Science and Math Carolina Biological Supply Company, 2003 Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Basics of Analytical Chemistry and Chemical Equilibria Brian M. Tissue, 2023-03-28 BASICS OF ANALYTICAL CHEMISTRY AND CHEMICAL EQUILIBRIA Familiarize yourself with the fundamentals of analytical chemistry with this easy to follow textbook Analytical chemistry is the study of chemical composition concerned with analyzing materials to discover their constituent substances the amounts in which these substances are present and more Since materials exist in different states and undergo reactions analytical chemistry is also concerned with chemical equilibria the state at which various reactants and substances will undergo no observable chemical change without outside stimulus This field has an immense range of practical applications in both industry and research and is a highly desirable area of expertise for the next generation of chemists Basics of Analytical Chemistry and Chemical Equilibria provides an introduction to this foundational subject ideal for specialized courses It introduces not only the core concepts of analytical chemistry but cultivates mastery of various instrumental methods by which students and researchers can undertake their own analyses Now updated to include the latest research and expanded coverage Basics of Analytical Chemistry and Chemical Equilibria promises to situate a new generation of readers in this growing field Readers of the second edition of Basics of Analytical Chemistry and Chemical Equilibria will also find A new chapter on structure determination Revised and expanded descriptions of chemical instrumentation You try it exercises throughout to further develop practical student knowledge Companion website of associated materials including end of chapter solutions spreadsheets for student use and more Basics of Analytical Chemistry and Chemical Equilibria is an ideal textbook for students in chemistry biochemistry and

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