

**Molar Mass Practice 1.8 – Answer Key**

**Objectives:** - Calculate molar mass using the periodic table

**Directions:** Find the molar mass of the following atoms, molecules, or formula units. Round masses to the hundredths place. Don't forget your units! You must show your work!

1. Zn

$$65.38 \text{ g/mol}$$

9.  $(\text{NH}_4)_2\text{S}$ 

$$2(14.01) + 8(1.01) + 32.07 = 96.17 \text{ g/mol}$$

2. Fe

$$55.85 \text{ g/mol}$$

10.  $\text{C}_2\text{H}_6\text{O}$ 

$$\text{C: } 2 \times 12.01 = 24.02$$

$$\text{H: } 6 \times 1.01 = 6.06$$

$$\text{O: } 1 \times 16.00 = 16.00$$

$$46.08 \text{ g/mol}$$

3.  $\text{CO}_2$ 

$$\text{C: } 1 \times 12.01 = 12.01$$

$$\text{O: } 2 \times 16.00 = 32.00$$

$$44.01 \text{ g/mol}$$

11.  $3\text{Mg}(\text{NO}_3)_2$ 

$$3\text{Mg: } 3 \times 24.31 = 72.93$$

$$\text{N: } 6 \times 14.01 = 84.06$$

$$\text{O: } 18 \times 16.00 = 288.00$$

$$445.00 \text{ g/mol}$$

4.  $\text{LiCl}$ 

$$\text{Li: } 1 \times 6.94 = 6.94$$

$$\text{Cl: } 1 \times 35.45 = 35.45$$

$$42.39 \text{ g/mol}$$

12.  $\text{H}_2\text{SO}_4$ 

$$2(1.01) + 32.07 + (4 \times 16.00) = 98.09 \text{ g/mol}$$

5.  $\text{O}_2$ 

$$\text{O: } 2 \times 16.00 = 32.00$$

$$32.00 \text{ g/mol}$$

13.  $\text{Mg}_3(\text{PO}_4)_2$ 

$$3\text{Mg: } 3 \times 24.31 = 72.93$$

$$\text{P: } 2 \times 30.97 = 61.94$$

$$\text{O: } 8 \times 16.00 = 128.00$$

$$262.87 \text{ g/mol}$$

6. Iron(III) oxide



$$2(55.85) + 3(16.00) = 159.7 \text{ g/mol}$$

14. Dihydrogen monoxide



$$2(1.01) + 16.00 = 18.02 \text{ g/mol}$$

7.  $\text{Hg}_2\text{BrO}_4$ 

$$459.86 \text{ g/mol}$$

15.  $\text{Zn}(\text{NO}_3)_2$ 

$$244.4 \text{ g/mol}$$

8. Lead(II) hydroxide



$$241.2 \text{ g/mol}$$

16. Aluminum sulfate



$$342 \text{ g/mol}$$

# Formula Or Molar Mass Answer Key

**SJ Ball**



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