

Chem 104A Study Guide

1. A gas is initially at 200 kPa pressure and 27°C. If the temperature is increased by 10 degrees Celsius, what will the new pressure be if volume remains constant?
2. The volume of a gas is 1.1 L and at 400 kPa. Assuming ideal behavior, what is the volume at the temperature of the gas?
3. A mixture that contained equal moles of O_2 , N_2 , CO_2 , and CH_4 (molar masses: 32, 28, 44, and 16 g/mol) was compressed from completely emptying 2.0 g of O_2 at 200 kPa. What pressure?
4. A mixture of 2.0 L of gas at a pressure of 10 kPa was at a temperature of 100°C. What will be the new pressure if the gas is compressed from 2.0 L to 1.0 L and the pressure is 10 kPa?
5. One liter mixture always in equilibrium that reacts at 100°C. The mixture is made up of 1.000 g of hydrogen gas at 100°C. The chemical equation for the reaction is: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$ at 100°C.

Gas Law Study Guide

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