

13. Complete the following table for the three experiments in Model 2.

	Experiment C	Experiment D	Experiment E
Independent Variable	# of particles	Temp.	Internal External Pressure
Dependent Variable	Volume	Volume	Volume
Controlled Variable(s)	Temp / I+E pressure	External + Internal Pressure / # of particles	Temp / # of particles

14. Provide a molecular level explanation for the increase in volume among the balloons in Experiment C. (How often and/or how hard are the molecules hitting the sides of the container?)

# of molecules ↑, so number of collisions ↑ + Volume along with it

collisions ↑ energy stays the same

15. Provide a molecular level explanation for the increase in volume among the balloons in Experiment D.

Temperature ↑ increasing the kinetic energy of the particles as well  
which ↑ collisions thus ↑ pressure

16. Provide a molecular level explanation for the increase in volume among the balloons in Experiment E.

if external forces decrease then it allows internal forces ↑  
thus ↑ volume

# Gas Variables Pogil Answer

**C. E. R. Bruce**

## **Gas Variables Pogil Answer:**

**Gases and Gas Laws** Hobsons Publishing, PLC,1983-01-01      **Gas Movements in the Long-period Variables** C. E. R. Bruce,1957      **An Introduction to the Gas Phase** Claire Vallance,2017-11-23 An Introduction to the Gas Phase is adapted from a set of lecture notes for a core first year lecture course in physical chemistry taught at the University of Oxford The book is intended to give a relatively concise introduction to the gas phase at a level suitable for any undergraduate scientist After defining the gas phase properties of gases such as temperature pressure and volume are discussed The relationships between these properties are explained at a molecular level and simple models are introduced that allow the various gas laws to be derived from first principles Finally the collisional behaviour of gases is used to explain a number of gas phase phenomena such as effusion diffusion and thermal conductivity

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