

# Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PV = nRT$ ", and the universal gas constant  $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.325 \text{ kPa} / 1 \text{ atm}$  to get  $R = 8.31 \text{ kPa} \cdot L / (K \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$
$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of  $87^\circ\text{C}$ , how many moles of gas do I have?

$$PV = nRT$$
$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$
$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of  $56^\circ\text{C}$ , what is the volume of the container that the gas is in?

$$PV = nRT$$
$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 328 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of  $67^\circ\text{C}$ , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$
$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 340 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$
$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

# Gas Law Questions And Answers

**J Rink**



## Gas Law Questions And Answers:

Chemistry All-in-One For Dummies (+ Chapter Quizzes Online) Christopher R. Hren, John T. Moore, Peter J. Mikulecky, 2022-11-23 Everything you need to crush chemistry with confidence Chemistry All in One For Dummies arms you with all the no nonsense how to content you'll need to pass your chemistry class with flying colors You'll find tons of practical examples and practice problems and you'll get access to an online quiz for every chapter Reinforce the concepts you learn in the classroom and beef up your understanding of all the chemistry topics covered in the standard curriculum Prepping for the AP Chemistry exam Dummies has your back with plenty of review before test day With clear definitions concise explanations and plenty of helpful information on everything from matter and molecules to moles and measurements Chemistry All in One For Dummies is a one stop resource for chem students of all valences Review all the topics covered in a full year high school chemistry course or one semester of college chemistry Understand atoms molecules and the periodic table of elements Master chemical equations solutions and states of matter Complete practice problems and end of chapter quizzes online Chemistry All In One For Dummies is perfect for students who need help with coursework or want to cram extra hard to ace that chem test

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*Gas World*, 1907

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## **Gas Law Questions And Answers** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Gas Law Questions And Answers**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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