

Gas Laws and Theory SCH3U

Name: _____

10 K 10 T 4 C 3 A Assessment of Learning (Product)

$$R = \frac{8.31 \text{ kPa} \cdot \text{L}}{\text{mol} \cdot \text{K}}$$

$$1 \text{ atm} = 101.325 \text{ kPa} = 760 \text{ torr} = 760 \text{ mmHg}$$

$$\text{SATP} = 25^\circ\text{C and } 100 \text{ kPa}$$

$$\text{STP} = 0^\circ\text{C and } 101.3 \text{ kPa}$$

Multiple Choice. Circle the most correct answer. (10K)

1. A Kelvin temperature that is equivalent to 500°C is:

- A) 227 K B) 773 K C) 1260 K D) -227 K

2. Convert the pressure 1.80 kPa to atm.

- A) 2 atm B) 115 atm C) 132 atm D) 245 atm E) None of the above

3. For any confined ideal gas, when the temperature is increased and the volume is decreased at the same time, the:

- a. molecules move more slowly b. average kinetic energy is decreased
c. number of collisions is decreased d. pressure stays constant e. pressure increases

4. The forces holding molecules together are strongest in a:

- A) solid B) liquid C) gas D) vapor

5. According to the kinetic-molecular theory, how does a gas expand?

- A) Its particles become larger. B) Collisions between particles become elastic.
C) Its temperature decreases. D) Its particles move greater distances apart.

6. What does the constant bombardment of gas molecules against the inside walls of a container produce?

- A) temperature B) density C) pressure D) diffusion

7. If the volume of a container of gas is reduced, while temperature is kept constant, what will happen to the pressure inside the container?

- A) The pressure will increase B) The pressure will not change
C) The pressure will decrease D) The pressure depends on the type of gas.

8. What happens to the pressure of a gas inside a container if the temperature of the gas decreases?

- A) The pressure increases. C) The pressure decreases.
B) The pressure does not change. D) The pressure cannot be predicted.

9. When the Kelvin temperature of an enclosed gas doubles, the particles of the gas _____

- A) move faster B) strike the walls of the container with less force
C) decrease in average kinetic energy D) decrease in volume

Gas Law Quiz With Answers

Charles William Gerstenberg

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