

GAS LAWS, GAS STOICHIOMETRY AND KINETIC MOLECULAR THEORY OF GASES

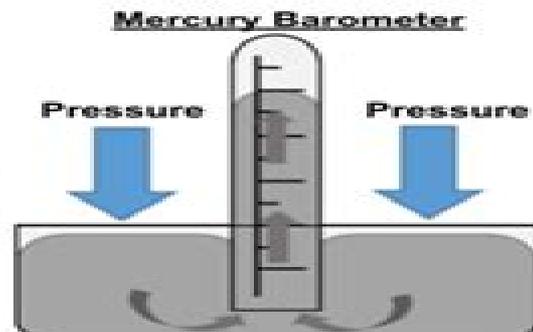
The entire universe is made up of matter including humans, animals, plants, and even the non-living things. There are three phases or states of matter, namely; solid, liquid, and gas. Moreover, gas behaves differently from solids and liquids due to differences in their molecular behavior. The movement of the gas is in random motion due to the vast empty space in a certain system since the distance between particles are far greater than the other two states of matter. Under some conditions of a gas sample, it can be defined in terms of its variables, temperature, volume, number of moles, and pressure.

In the gaseous phase, molecules or the particles collide randomly against other molecules and against their container or their system. This random collision results changes in momentum which give way to the one of the properties of gases called pressure. Pressure, as one of the variables, is defined as the force exerted per unit area. It refers to the force exerted by colliding molecules per unit area of container walls and is mathematically described as $P = F/A$.

Units of Pressure

What makes pressure quantifiable? How is pressure measured and expressed? The following details are some of the units that can be used for pressure:

- Standard Atmosphere (atm)** □ It is commonly and widely used unit for pressure in chemistry. The 1 atm is equal to the pressure that supports a column of mercury which is exactly 760 mmHg.
- Torr (or mmHg)** □ mmHg means millimeter of mercury, where it represents the pressure exerted by a column of a mercury which exactly equals to atmosphere. The unit mmHg is called torr, named after the Italian scientist *Evangelista Torricelli*, who also invented *Barometer*, a device used to measure the atmospheric pressure.
- Pounds per square inch (psi)** □ If we say that gas exerts a pressure of 20 psi, it means the pressure on the wall of the gas container is 20 pounds or the force per square inch of the unit area. The atmospheric pressure at sea level is 14.7 psi.
- kilopascal(kPa)** □ A kilopascal is equal to 1000 pascals (Pa). It is the SI unit for pressure.



Conversion Factor:

$$1 \text{ atm} = 760 \text{ mmHg} = 760 \text{ torr} = 101.3 \text{ kPa} = 14.7 \text{ psi}$$
$$1 \text{ kPa} = 1000 \text{ Pa}$$

Air is all around us, thus, gases are always involved in our daily activities, from breathing down to automobile or bicycle tires, balloons, and even lifeboats and vests. Life won't be possible without this life-sustaining gas found in the atmosphere.

Scientists have always been curious about how gases behave and how they are different compared to other states of matter. Investigations and experiments on the behavior of gases lead

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JR Anderson



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