

Boyle's Law	Charles' Law	Guy-Lassac's Law	Combined Gas Law
For a given mass of gas at constant temperature, the volume of a gas varies inversely with pressure	The volume of a fixed mass of gas is directly proportional to its Kelvin temperature if the pressure is kept constant.	The pressure of a gas is directly proportional to the Kelvin temperature if the volume is kept constant.	Combines Boyle's, Charles', and the Temperature-Pressure relationship into one equation. Each of these laws can be derived from this law.
$PV = k$ $P_1V_1 = P_2V_2$	$\frac{V}{T} = k$ $V_1T_2 = V_2T_1$ $\frac{V_1}{T_1} = \frac{V_2}{T_2}$	$\frac{P}{T} = k$ $P_1T_2 = P_2T_1$ $\frac{P_1}{T_1} = \frac{P_2}{T_2}$	$\frac{PV}{T} = k$ $V_1P_1T_2 = V_2P_2T_1$ $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

Dalton's Law	Ideal Gas Law	Graham's Law
At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the pressures exerted by each gas.	The Ideal Gas Law relates the pressure, temperature, volume, and mass of a gas through the gas constant "R".	The rate of effusion/diffusion of two gases (A and B) are inversely proportional to the square roots of their formula masses. [It can be a ratio of molecular speeds, effusion /diffusion times, distance traveled by molecules, or amount of gas effused]
$P_{total} = P_1 + P_2 + P_3 + \dots P_n$	$PV = nRT$	$\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

Abbreviations	Standard Conditions
atm = atmosphere mm Hg = millimeters of mercury torr = another name for mm Hg Pa = Pascal kPa = kilopascal K = Kelvin °C = degrees Celsius	0°C = 273 K 1.00 atm = 760.0 mm Hg = 76 cm Hg = 101.325 kPa = 101, 325 Pa = 29.9 in Hg
Conversions	Gas Law's Equation Symbols
$K = ^\circ C + 273$ $F = 1.8C + 32$ $C = \frac{F - 32}{1.8}$ 1 cm ³ (cubic centimeter) = 1 mL (milliliter) 1 dm ³ (cubic decimeter) = 1 L (liter) = 1000 mL	Subscript (1) = old condition or initial condition Subscript (2) = new condition or final condition Temperature must be in Kelvins n = number of moles = grams/Molar mass R = 8.31 L·kPa/ mol·K = 0.0821 L·atm/mol·K = 62.4 L·Torr/mol·K You must have a common set of units in the problem

Gas Laws Packet 1 Answer Key

Rachel S Tattersall



Gas Laws Packet 1 Answer Key:

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In some sort of inundated with displays and the cacophony of fast connection, the profound power and mental resonance of verbal art frequently disappear into obscurity, eclipsed by the constant barrage of noise and distractions. However, nestled within the lyrical pages of **Gas Laws Packet 1 Answer Key**, a fascinating function of literary brilliance that pulses with organic emotions, lies an unforgettable journey waiting to be embarked upon. Published by way of a virtuoso wordsmith, that interesting opus guides visitors on an emotional odyssey, softly exposing the latent potential and profound influence embedded within the complicated internet of language. Within the heart-wrenching expanse of this evocative analysis, we shall embark upon an introspective exploration of the book is central styles, dissect their fascinating publishing fashion, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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