

## Gas Laws Practice - Combined, Individual, and Ideal Gas Laws (Answers on next page)

Ideal Gas Law:

$$PV = nRT$$

$$R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$$

$$R = 62.4 \frac{\text{L} \cdot \text{mmHg}}{\text{mol} \cdot \text{K}}$$

$$R = 8.314 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}}$$

Combined Law:

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Important Conversions:

$$\text{K} = ^\circ\text{C} + 273$$

$$1 \text{ atm} = 760 \text{ mmHg} = 760 \text{ Torr} = 101.3 \text{ kPa}$$

*Note: Torr is just another pressure unit that is equivalent to mmHg.*

- It is not safe to put aerosol canisters in a campfire because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2.0 moles of gas, and the campfire temperature is 1400.<sup>o</sup> C, what is the pressure in mmHg inside the canister?
- A sample of nitrogen gas is cooled from 1200.<sup>o</sup> C. What is its final temperature (in K) if it compresses from 12.4 L to 2.50 liters?
- Zach has added some volume (in milliliters) of air to a balloon at sea level (1.0 atm). If I take the balloon with me to Denver where the air pressure is 0.85 atm, the volume of the balloon is 15.0 L. What was the original volume of the balloon?
- How many moles of gas are in a 30.0 liter scuba canister if the temperature of the canister is 300. K and the pressure is 20300 kPa?
- A small research submarine with a volume of  $1.2 \times 10^5$  L has an internal pressure of 1.0 atm and an internal temperature of 15<sup>o</sup> C. If the submarine descends to a depth where the pressure is 150 atm and the temperature is 3.0<sup>o</sup> C, what would be the new volume (in L) of the gas if the hull of the submarine is crushed?
- If 35.0 grams of helium are held at 115 kPa and 120 liters, what is the temperature (in <sup>o</sup>C)?
- What pressure (in Torr) is required to compress 196.0 liters of ammonia (NH<sub>3</sub>) at 1.00 atmosphere into a cylinder whose volume is 26.0 liters? Assume temperature stays constant.
- My car has a constant internal volume of 2,600 liters. If the sun heats my car from a temperature of 20.<sup>o</sup>C to a temperature of 55 <sup>o</sup>C, what will the pressure inside my car be if the pressure was initially 760 mmHg?
- How many grams of gas are initially in my car if the gas is oxygen (O<sub>2</sub>) in question #8?
- Consider a closed container of gas.
  - Why does the pressure of the gas increase if more gas is added to the container if all other variables are kept constant?
  - If the volume of the container is compressed, why does pressure of the gas increase if all other variables are kept constant?
  - If the volume of the container is compressed, why does temperature of the gas decrease if all other variables are kept constant?
  - If the gas is cooled, why does pressure of the gas decrease if all other variables are kept constant?

# Gas Laws Practice Calculations Answers Unit Lesson 01

**Daniel C. Tofan**



## Gas Laws Practice Calculations Answers Unit Lesson 01:

**General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

**Resources in Education**, 1973

**Fundamentals of Physics, Volume 1** David Halliday, Robert Resnick, Jearl Walker, 2017-12-11 The first volume of a two volume text that helps students understand physics concepts and scientific problem solving Volume 1 of the Fundamentals of Physics 11th Edition helps students embark on an understanding of physics This loose leaf text covers a full range of topics including measurement vectors motion and force It also discusses energy rotation equilibrium gravitation and oscillations as well temperature and heat The First and Second Law of Thermodynamics are presented as is the Kinetic Theory of Gases The text problems questions and provided solutions guide students in improving their problem solving skills

Introductory Chemistry Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students

*Research in Education*, 1973

*Chemical Principles* Monty J. Montjar, 1990

*CHEMISTRY SILBERBERG*, 2003

*Essentials of Introductory Chemistry* Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids

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