

Gas Laws

Avogadro's Law:

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

Boyle's Law:

$$V \propto \frac{1}{P} \quad P_1V_1 = P_2V_2$$

Charles' Law:

$$V \propto T \quad \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Gay-Lussac's Law:

$$P \propto T \quad \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

Combined Gas Law:

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

Dalton's Law:

$$P_{\text{total}} = P_1 + P_2 + P_3 + \dots$$

Ideal Gas Law:

$$PV = nRT \quad \left(R = \frac{0.08206 \text{ L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \right)$$

Conversions:

$$\begin{aligned} 1 \text{ atm} \\ &= 760 \text{ mmHg} \\ &= 760 \text{ torrs} \\ &= 101325 \text{ Pa} \\ &= 101.35 \text{ kPa} \end{aligned}$$

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