

KEY

Gas Stoichiometry Worksheet

Directions: Use the gas laws we have learned to solve each of the following problems. Each of the chemical equations must first be balanced. Show all your work for credit.

1. When calcium carbonate is heated strongly, carbon dioxide gas is released according to the following equation:



What volume of $\text{CO}_2(g)$, measured at STP, is produced if 15.2 grams of $\text{CaCO}_3(s)$ is heated?

$$\frac{15.2 \text{ g CaCO}_3}{100.1 \text{ g}} \times \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CaCO}_3} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol CaCO}_3} \times \frac{22.4 \text{ L CO}_2}{1 \text{ mol CO}_2} = 3.40 \text{ L CO}_2$$

2. The synthesis of sodium chloride occurs according to the reaction:



What volume of chlorine at STP is necessary for the complete reaction of 4.81 grams of sodium metal?

$$\frac{4.81 \text{ g Na}}{23.0 \text{ g}} \times \frac{1 \text{ mol Na}}{1 \text{ mol Na}} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \times \frac{22.4 \text{ L Cl}_2}{1 \text{ mol Cl}_2} = 2.34 \text{ L Cl}_2$$

3. Potassium permanganate is produced commercially by the this reaction.



What volume of chlorine gas at STP would be required to produce 10.0 grams of KMnO_4 ?

$$\frac{10.0 \text{ g}}{158.0 \text{ g}} \times \frac{1 \text{ mol KMnO}_4}{1 \text{ mol KMnO}_4} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol KMnO}_4} \times \frac{22.4 \text{ L}}{1 \text{ mol}} = 7.09 \text{ liters}$$

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