

Gastrointestinal System

ONE STOP DOC



- the core questions and answers you need
- the key concepts you should understand
- the key facts you should know

in one concise integrated revision guide



Miruna Canagaratnam
Editorial Advisor: Richard Naffallo

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Miruna Canagatnam



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net ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} + \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g} \rightarrow \text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions
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web the six NO_3^- ions and the six Na^+ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3\text{PO}_4\text{2 s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3 \text{aq} + \text{KCl} \text{aq} \rightarrow \text{AgCl} \text{s} + \text{KNO}_3 \text{aq}$ total ionic $\text{Ag}^+ \text{aq} + \text{NO}_3^- \text{aq} + \text{K}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s} + \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$ net ionic $\text{Ag}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{cl aq} + \text{cu}_2 \text{aq} \rightarrow \text{cucl}_2 \text{s} + 2\text{cl aq}$ $\text{cu}_2 \text{aq} + \text{cu}_2 \text{aq} \rightarrow \text{cucl}_2 \text{s}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length