

## ① Atoms



Particle	P	N	E
Charge	+1	0	-1
Mass	1	1	0

Isotopes - A different atomic form of the same element, has the same number of protons, different number of neutrons.

<sup>12</sup> C	<sup>13</sup> C
<sup>14</sup> C	
<sup>15</sup> C	
<sup>16</sup> C	
<sup>17</sup> C	
<sup>18</sup> C	

$$\text{Relative atomic mass (Ar)} = \frac{\text{Sum of (isotope abundance} \times \text{isotope mass no.)}}{\text{Sum of abundances of all the isotopes}}$$

## ② Compounds

A substance made of atoms of at least two different elements, chemically bonded together.

## ③ Elements

All atoms of an element have the same number of protons.

Different elements have atoms with different numbers of protons.

## ④ Chemical Equations

During a chemical reaction bonds are broken, atoms rearrange, new bonds form



Balance:



Topic 1a -  
Atoms, Elements,  
Compounds &  
Mixtures

## ⑤ Electronic Configuration

Electrons occupy shells (energy levels)  
Electron shells with lowest energy is filled first

Shell	Max no. of electrons
1	2
2	8
3	8

## ⑥ Separating Techniques

There are a number of physical techniques used to separate mixtures of solids, liquids & gases.

Technique	What is separated
Paper Chromatography	Mix of different liquids
Filtration	Insoluble solids from liquid
Evaporation & Crystallisation	Soluble solids from liquid
Distillation	Mixture of liquids depending on boiling point

## ⑦ The History of the Atom

Year	Scientist	Model	Structure
1803	John Dalton	Solid sphere	
1904	J. J. Thompson	Plum pudding	
1911	Ernest Rutherford	Nuclear	
1913	Niels Bohr	Planetary	

# Gcse Science Chemistry Revision Guide

**Karin Nielsen-Saines**



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