

## Naming Acids

What makes an acid unique and how are acids named?

### Why?

A variety of acids are used in foods, industry, and research. Acids are covalently bonded molecules, but when they are put into water they produce ions. One of the ions produced is always  $H^+$ , which immediately combines with a water molecule to form the hydronium ion ( $H_3O^+$ ). The  $H_3O^+$  ion is what defines the acidic properties of a substance. Because of their special classification, acids have a naming system different from ionic or other molecular (covalent) compounds.

### Model 1 – Binary Acids

Acid	Name of acid in aqueous solution	Cation (+)*	Anion (-)
HCl	Hydrochloric acid	$H_3O^+$	$Cl^-$
HBr	Hydrobromic acid	$H_3O^+$	$Br^-$
$H_2S$	Hydrosulfuric acid	$2H_3O^+$	$S^{2-}$
HF	Hydrofluoric acid	$H_3O^+$	$F^-$

\*Hydrogen ions ( $H^+$ ) join with water molecules in solution to form hydronium ions,  $H_3O^+$ .

1. Complete the cation and anion columns of the table in Model 1. Be careful to show the charges on the ions.
2. Why does hydrosulfuric acid contain two hydrogens?

*Two hydrogens with a +1 charge are required to balance the 2- charge of the sulfide anion.*

3. Look at the formulas and names of the binary acids in Model 1.
  - a. What prefix is used at the beginning of the name for all binary acids?
  - b. What suffix is used at the end of the name for all binary acids?

*Binary acids begin with the prefix "hydro-."*

*The names of binary acids end with "-ic," followed by the word "acid."*

4. The prefix "bi" means "two." Propose a reason that the acids in Model 1 are all referred to as "binary" acids.

*All binary acids have two different elements or atoms in them.*



5. Write a rule for naming binary acids.

*All binary acids begin with the prefix hydro-, which precedes the name of any nonmetal other than hydrogen, in the compound. The name of this nonmetal is modified to end with the suffix -ic.*

*Example: hydro\_\_\_\_\_ic acid.*



# Naming Acids Packet Answer Key

**Sir Tarun Rupani**



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## **Table of Contents Naming Acids Packet Answer Key**

1. Understanding the eBook Naming Acids Packet Answer Key
  - The Rise of Digital Reading Naming Acids Packet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Naming Acids Packet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Naming Acids Packet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Naming Acids Packet Answer Key
  - Personalized Recommendations
  - Naming Acids Packet Answer Key User Reviews and Ratings
  - Naming Acids Packet Answer Key and Bestseller Lists

5. Accessing Naming Acids Packet Answer Key Free and Paid eBooks
  - Naming Acids Packet Answer Key Public Domain eBooks
  - Naming Acids Packet Answer Key eBook Subscription Services
  - Naming Acids Packet Answer Key Budget-Friendly Options
6. Navigating Naming Acids Packet Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Naming Acids Packet Answer Key Compatibility with Devices
  - Naming Acids Packet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Naming Acids Packet Answer Key
  - Highlighting and Note-Taking Naming Acids Packet Answer Key
  - Interactive Elements Naming Acids Packet Answer Key
8. Staying Engaged with Naming Acids Packet Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Naming Acids Packet Answer Key
9. Balancing eBooks and Physical Books Naming Acids Packet Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Naming Acids Packet Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Naming Acids Packet Answer Key
  - Setting Reading Goals Naming Acids Packet Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Naming Acids Packet Answer Key
  - Fact-Checking eBook Content of Naming Acids Packet Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- 
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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