

1. A gas has a pressure of 0.370 atm at 55.0 °C. What is the pressure at standard temperature?

$$P_2 = P_1 \frac{T_2}{T_1} = \frac{(0.370 \text{ atm})(273 \text{ K})}{328 \text{ K}} = 0.313 \text{ atm}$$

2. If 2.00 moles of gas occupies 4.00L at STP. How much of the same gas will occupy 3.00L at STP?

$$n_2 = n_1 \frac{V_2}{V_1} = \frac{(2.00 \text{ moles})(3.00 \text{ L})}{(4.00 \text{ L})} = 1.50 \text{ mole}$$

3. I've got a car with an internal volume of 12,000 L. If I drive my car into the river and it implodes, what will be the volume of the gas when the pressure goes from 1.0 atm to 1.4 atm?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(1.0 \text{ atm})(12,000 \text{ L})}{(1.4 \text{ atm})} = 8571 \text{ L}$$

4. What's the partial pressure of carbon dioxide in a container that holds 5.00 moles of carbon dioxide, 3.00 moles of nitrogen, and 1.00 mole of hydrogen and has a total pressure of 1.05 atm?

$$P_{\text{CO}_2} = P_1 \times \frac{\text{mole CO}_2}{\text{Total moles}} = (1.05 \text{ atm}) \times \frac{(5.00 \text{ moles})}{(9.00 \text{ moles})} = 0.583 \text{ atm}$$

5. I have 130 liters of gas in a piston at a temperature of 250. °C. If I cool the gas until the volume decreases to 85 liters, what will temperature of the gas be?

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(85 \text{ L})(250 \text{ K})}{130 \text{ L}} = 162 \text{ K} = -111 \text{ °C}$$

6. If a gas in a closed container is pressurized from 15.0 atmospheres to 16.0 atmospheres and its original temperature was 25.0 °C, what would the final temperature of the gas be?

$$T_2 = \frac{P_2 T_1}{P_1} = \frac{(16.0 \text{ atm})(298 \text{ K})}{(15.0 \text{ atm})} = 318 \text{ K}$$

7. To compress nitrogen at 1 atm from 750 mL to 500 mL, what must the new pressure be if the temperature is kept constant?

$$P_2 = \frac{V_1 P_1}{V_2} = \frac{(750 \text{ mL})(1 \text{ atm})}{(500 \text{ mL})} = 1.5 \text{ atm}$$

8. Calcium carbonate decomposes at 1200. °C to form carbon dioxide and calcium oxide. If 25 liters of carbon dioxide are collected at 1200. °C, what will the volume of this gas be after it cools to 25 °C?

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

9. 5.00 L of a gas is known to contain 0.265 moles. If this amount of gas is increased to 1.80 moles, what new volume will result (at an unchanged temperature and pressure)?

$$V_2 = \frac{V_1 n_2}{n_1} = \frac{(5.00 \text{ L})(1.80 \text{ moles})}{(0.265 \text{ moles})} = 337 \text{ L}$$

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