

# Gas Variables

How are the variables that describe a gas related?

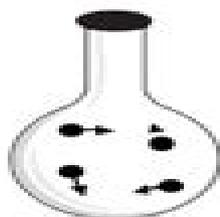
## Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm?

What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar<sup>®</sup>, which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

## Model 1 - Gases in a Nonflexible Container

### Experiment A (Adding more gas)



A1

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 1 atm  
Temperature = 200 K



A2

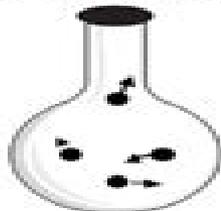
Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 2 atm  
Temperature = 200 K



A3

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 3 atm  
Temperature = 200 K

### Experiment B (Heating the gas)



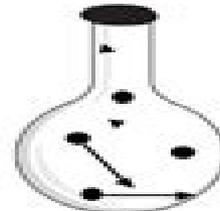
B1

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 1 atm



B2

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 2 atm



B3

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 3 atm

# Gas Variables Packet

**Thomas P. Espinola**

A red circular graphic with a gradient, appearing as a semi-circle on the right side of the slide, partially overlapping the grey bar.

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