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PRECIPITATION REACTION LAB

Introduction:

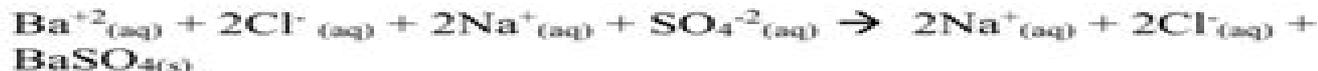
The majority of ionic solids like salts (KCl and NaCl) are soluble in water because the polar water molecules surround the individual ions of the salt. Those that do not dissolve and go into solution form solid products called precipitates. These precipitates have many colors and often help scientists identify what the precipitate is present. In meteorology the term precipitation refers to meteorological phenomena such as rain or snow. Precipitation in chemical solutions occurs when two chemicals react to form a product that is insoluble in water and falls out of solution like rain or snow. A precipitate is a solid substance that separates from solution during a chemical reaction. A precipitate can be identified by the cloudy, milky, gelatinous, or grainy appearance it gives to the mixture.

A barium sulfate precipitate can be produced by the reaction of barium chloride and sodium sulfate. A chemical equation to describe the reaction is written and balanced like this:



Barium sulfate, $\text{BaSO}_4 (\text{s})$, is a common precipitate used as an X-ray contrast medium because it is insoluble in water and opaque to x-rays. Typically a patient drinks an aqueous slurry of barium sulfate just before he is x-rayed. The precipitate coats his stomach and intestines. These organs then show up on the x-ray film in vivid contrast, aiding the doctor's diagnosis. Notice the reaction that forms BaSO_4 is a double-replacement reaction in which the cations and anions of the reactants trade partners to form the products.

To write a total ionic equation, rewrite all aqueous substances as their component ions and keep all solid substances unchanged.



Cancel all components of the reaction that are identical on both the reactant and product side of the reaction. These are called spectator ions. Write the net ionic equation by rewriting the reaction equation including only those reactants and products that change in the reaction.

Investigating Precipitation Reactions Lab Answers

Dennis D. Staley, Edward L. Waterman



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