

Chemistry Benchmark II Study Guide

- Write the formula for the following:
 - Iron (II) oxide
 - calcium phosphate
 - potassium bromide
 - ammonium phosphate
 - copper (II) bromide
 - magnesium hydroxide
 - nitrogen dioxide
 - titanium (IV) oxide
 - xenon hexafluoride
- What is a binary compound? What is a ternary compound?
- Write formulas for
 - hydroiodic acid
 - carbonic acid
 - phosphorous acid
 - sulfuric acid
 - hydrosulfuric acid
 - hydrochloric acid
 - chloric acid
 - chlorous acid
 - hypochlorous acid
 - perchloric acid
- Name these:
 - NaCl
 - KI
 - CsBr
 - Rb₂O
 - Sr₂
 - CuCl
 - HgO
 - Fe₂O₃
 - MnO₂
 - PbCl₂
 - CuI
 - SnBr₄
 - BF₃
 - NO
 - N₂O₅
 - CCl₄
 - P₂O₅
 - N₂O₃
- Find the oxidation number assigned to:
 - S in H₂SO₄
 - P in PCl₅
 - S in SO₂
- Name these.
 - NO₂
 - N₂O
 - SO₂
- How many valence electrons do elements have in the following groups?
 - alkali metals
 - alkaline earth metals
 - halogens
 - noble gases
- Write the complete electron configuration, noble gas configuration, Lewis dot structure, and orbital notation for:
 - Sr
 - Br
 - Na
 - K
 - N
 - Cl
- In each group, identify the atom with the largest atomic radius.
 - Sn, Xe, Rb, Sr
 - Rn, He, Xe, Kr
 - Pb, Ba, Cs, At
- In each group, identify the atom with the highest ionization energy.
 - Cs, K, Li
 - Ba, Sr, Ca
 - I, Br, Cl
 - Mg, Si, S
- Define the following: ionic bond, covalent bond, polar covalent bond, nonpolar covalent bond
- Define electronegativity and describe the trend across a period and down a group. What type of element has the highest electronegativity? What type has the lowest electronegativity?
- Identify the type of bond in each:
 - K-N
 - Cs-O
 - C-N
 - O-F
 - Na-Cl
 - Cl-F
 - Cl-O
 - Ca-Cl
- Who first arranged the periodic table in order of increasing atomic mass?
- Draw Lewis structures for the following and predict the molecular shape.
 - H₂S
 - PH₃
 - O₂
 - sulfate ion
 - SO₂
 - BF₃
- Calculate the wavelength and energy of 96.6MHz radio waves.
- What is the frequency and energy of electromagnetic radiation having a wavelength of $3.33 \times 10^{-6}\text{m}$?
- Predict the ions that the following elements might form:
 - F
 - K
 - I
 - Ga
 - Fe (trick question)
 - P
 - Ag

General Chemistry Benchmark 2 Study Guide

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Table of Contents General Chemistry Benchmark 2 Study Guide

1. Understanding the eBook General Chemistry Benchmark 2 Study Guide
 - The Rise of Digital Reading General Chemistry Benchmark 2 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Benchmark 2 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a General Chemistry Benchmark 2 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Benchmark 2 Study Guide
 - Personalized Recommendations
 - General Chemistry Benchmark 2 Study Guide User Reviews and Ratings
 - General Chemistry Benchmark 2 Study Guide and Bestseller Lists

5. Accessing General Chemistry Benchmark 2 Study Guide Free and Paid eBooks
 - General Chemistry Benchmark 2 Study Guide Public Domain eBooks
 - General Chemistry Benchmark 2 Study Guide eBook Subscription Services
 - General Chemistry Benchmark 2 Study Guide Budget-Friendly Options
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Benchmark 2 Study Guide
 - Highlighting and Note-Taking General Chemistry Benchmark 2 Study Guide
 - Interactive Elements General Chemistry Benchmark 2 Study Guide
8. Staying Engaged with General Chemistry Benchmark 2 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Benchmark 2 Study Guide
9. Balancing eBooks and Physical Books General Chemistry Benchmark 2 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Benchmark 2 Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Benchmark 2 Study Guide
 - Setting Reading Goals General Chemistry Benchmark 2 Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Benchmark 2 Study Guide
 - Fact-Checking eBook Content of General Chemistry Benchmark 2 Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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