

Name Key

1. Hydrogen,  $H_2$ , is used as rocket fuel. The hydrogen is burned in oxygen,  $O_2$ , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen,  $H_2$  is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$99.2 \text{ mol } H_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol } H_2} = -24,006 \text{ kJ}$$

-24,006 kJ  
~~24,006 kJ~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is  $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$ . How much heat must be added to 5.2 g of iron at  $25^\circ\text{C}$  to raise its temperature to  $225^\circ\text{C}$ ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine,  $N_2H_4$ , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



Reverse  
 multiply quantity

# General Chemistry Practice Problems With Answers

**Kenneth W. Whitten, Raymond E. Davis**

## **General Chemistry Practice Problems With Answers:**

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific, detailed, step-by-step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. **Key Features:** The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step-by-step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid-base reactions, which often persists through working with acid-base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

**General Organic and Biological Chemistry** Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health-related fields such as nursing, dental hygiene, nutrition, medical technology, and occupational therapy. It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry. Students need have no previous background in chemistry but should possess basic math skills. The text features numerous helpful problems and learning features.

**Student Solutions Manual** Darrell Ebbing, 2005-03-17 Provides solutions to odd-numbered Practice Problems, General Problems, and Cumulative Skills Problems, plus answers to Review Questions.

**General Chemistry, Reactions First** Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease. Using a friendly approach, the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry.

Problem Solving in General Chemistry Ronald Ragsdale, 1991

Survival Guide to Organic Chemistry Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19 The Survival Guide to Organic Chemistry: Bridging the Gap from General Chemistry enables organic chemistry students to bridge the gap between general chemistry and organic chemistry. It makes sense of the myriad of in-depth concepts of organic chemistry without overwhelming them in the necessary detail often given in a complete organic chemistry text. Here, the topics covered span the entire standard organic chemistry curriculum. The authors describe subjects which require further explanation, offer alternate viewpoints for understanding, and provide hands-on practical problems and solutions to help master the material. This text ultimately allows students to apply key ideas from their general chemistry curriculum to key concepts in organic chemistry.

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*Introduction to General, Organic, and Biochemistry* Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

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