

---

$$2.1 \quad p = pRT = (1.2)(1.01 \times 10^5) / (287)(300)$$

$$p = 1.41 \text{ kg/m}^3$$

$$v = 1/\rho = 1/1.41 = 0.71 \text{ m}^3/\text{kg}$$

---

2.2 Mean kinetic energy of each atom =

$$\frac{3}{2} kT = \frac{3}{2} (1.38 \times 10^{-23})(500) = 1.035 \times 10^{-20} \text{ J}$$

One kg-mole, which has a mass of 4 kg, has  $6.02 \times 10^{23}$  atoms. Hence 1 kg has  $\frac{1}{4}$  ( $6.02 \times 10^{23}$ ) =  $1.505 \times 10^{23}$  atoms.

Total internal energy = (energy per atom)(number of atoms)

$$= (1.035 \times 10^{-20})(1.505 \times 10^{23}) = 1.558 \times 10^3 \text{ J}$$

---

$$2.3 \quad p = \frac{P}{RT} = \frac{2116}{(1716)(460 + 59)} = 0.00237 \frac{\text{slug}}{\text{ft}^3}$$

$$\text{Volume of the room} = (20)(15)(8) = 2400 \text{ ft}^3$$

$$\text{Total mass in the room} = (2400)(0.00237) = 5.688 \text{ slug}$$

$$\text{Weight} = (5.688)(32.2) = 183 \text{ lb}$$

---

$$2.4 \quad p = \frac{P}{RT} = \frac{2116}{(1716)(460 + 10)} = 0.00274 \frac{\text{slug}}{\text{ft}^3}$$

Since the volume of the room is the same, we can simply compare densities between the two problems.

# Introduction To Flight 6th Solutions Manual

**K Payea**



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