

## Section 12.1

### OBJECTIVE 1

Write complete balanced molecular equations that are indicated.

1.  $\text{H}_2\text{SO}_4 + \text{NaOH}$ 
  - (a) Write the balanced molecular equation for the reaction.
  - (b) Write the complete ionic equation for the reaction.
  - (c) Write the net ionic equation for the reaction.
2.  $\text{H}_2\text{SO}_4 + \text{Ca(OH)}_2$ 
  - (a) Write the balanced molecular equation for the reaction.
  - (b) Write the complete ionic equation for the reaction.
  - (c) Write the net ionic equation for the reaction.
3.  $\text{H}_2\text{SO}_4 + \text{CaCl}_2$ 
  - (a) Write the balanced molecular equation for the reaction.
  - (b) Write the complete ionic equation for the reaction.
  - (c) Write the net ionic equation for the reaction.
4.  $\text{H}_2\text{SO}_4 + \text{BaCl}_2$ 
  - (a) Write the balanced molecular equation for the reaction.
  - (b) Write the complete ionic equation for the reaction.
  - (c) Write the net ionic equation for the reaction.

5. Write the net ionic equation for the following molecular equation:



a. Balance the equation:



b. Write the molecular equation for the net ionic equation.

c. Write the net ionic equation for the reaction between  $\text{MgSO}_4$  and  $\text{BaCl}_2$ .

d. Write the net ionic equation for the reaction between  $\text{MgSO}_4$  and  $\text{BaCl}_2$ .

e. Write the net ionic equation for the reaction between  $\text{MgSO}_4$  and  $\text{BaCl}_2$ .

# Introduction To Stoichiometry Section Review 9 1 Worksheet

**G Orfield**



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