

pH Properties of Buffer Solutions

Inquiry Guidance and AP[®] Chemistry Curriculum Alignment



Introduction

The physiological role of buffers within cells and in consumer products highlights the ability of buffers to resist changes in pH. Buffers provide an essential acid-base balancing act—in foods and drugs, consumer products, lakes and streams, and even living cells. All biological cells depend on the properties of buffers, as does the essential function of the respiratory system, breathing, which must be regulated within a very narrow pH range. What are buffers made of and how do they work?

Opportunities for Inquiry

Buffers represent an integrated concept bridging many big ideas. Understanding buffers requires students to synthesize knowledge and apply reasoning skills from chemical reactions, stoichiometry, and acid-base equilibrium. The myriad applications of buffers offer many interesting opportunities to study the composition and mechanism of action of buffers.

- Promote inquiry into buffers using a colorful demonstration. Dissolve Alka-Seltzer[®] tablets in water, add universal indicator, and observe the color and pH changes when 1 M HCl and 1 M NaOH are added. Ask students to design similar experiments to identify the potential buffering characteristics of other consumer products, including antacids, deodorants and conditioners, mouthwashes, beverages, contact lens solution, etc.
- Acetate buffers provide a convenient introductory activity to study how the pH of a buffer depends on the relative amounts of a weak acid and its conjugate base. Extend the baseline activity by asking students to select weak acid-conjugate base combinations that will be effective at different pH values across typical physiological or aqueous chemistry. Make it a challenge! Identify the ratios and concentrations at each pH that will have the highest capacity with respect to added acid or added base. *Hint:* They are not the same!
- Challenge students to prepare a set of buffer standards for an indicator color series that can be used for narrow-range pH analysis (a 0.2 pH units). An example is a series of phosphate buffers containing various ratios of NaH_2PO_4 and Na_2HPO_4 solutions with bromothymol blue (pH = 6.0–7.6).
- Forensic analysis of DNA by electrophoresis requires a buffer that will keep the charge on the DNA molecules relatively constant so that their migration in an electric field will depend only on their size. Research, prepare and study the pH profiles of electrophoresis buffers.
- Most enzymes have an optimum pH value for maximum activity. How are buffer solutions used to determine the optimum pH for a common enzyme such as amylase, catalase or pectinase?

Alignment with AP Chemistry Curriculum Framework—Big Ideas 3 and 6

Enduring Understandings and Essential Knowledge

Chemical changes are represented by a balanced chemical equation that identifies the ratios with which reactants react and products form. (Enduring Understanding 3A)

3A2: Quantitative information can be derived from stoichiometric calculations that utilize the mole ratios from the balanced chemical equations. The role of stoichiometry in real-world applications is important to note, so that it does not seem to simply be an exercise done only by chemists.

Chemical reactions can be classified by considering what the reactant are, what the product are, or how they change from one into the other. Classes of chemical reactions include synthesis, decomposition, acid-base, and oxidation-reduction reactions. (Enduring Understanding 3B)

3B2: In a neutralization reaction, protons are transferred from an acid to a base.

Chemical equilibrium plays an important role in acid-base chemistry and in solubility. (Enduring Understanding 6C)

6C1: Chemical equilibrium reasoning can be used to describe the proton transfer reactions of acid-base chemistry.

6C2: The pH is an important characteristic of aqueous solutions that can be controlled with buffers. Comparing pH to pK_a allows one to determine the protonation state of a molecule with a labile proton.

Flinn Ph Properties Of Buffer Solutions Answer Key

JR Anderson



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The Journal of Biological Chemistry ,1922 Vols 3 include the society s Proceedings 1907 Metals Abstracts ,1986
Dissertation Abstracts International ,1970 **Cumulated Index Medicus** ,1966 **Buffer Solutions** Professor Rob Beynon,J Easterby,2004-03-01 An indispensable guide to buffers and to understanding the principles behind their use Helps the user to avoid common errors in preparing buffers and their solutions A must for researchers in the biological sciences this valuable book takes the time to explain something often taken for granted buffers used in experiments It answers the common questions such as which buffer should I choose What about the temperature effects What about ionic strength Why is the buffer with the biggest temperature variation used in PCR It provides even the most experienced researchers with the means to understand the fundamental principles behind their preparation and use an indispensable guide essential for everyone using buffers **Buffers for pH and Metal Ion Control** D. Perrin,2012-12-06 This book is intended as a practical manual for chemists biologists and others whose work requires the use of pH or metal ion buffers Much information on buffers is scattered throughout the literature and it has been our endeavour to select data and instructions likely to be helpful in the choice of suitable buffer substances and for the preparation of appropriate solutions For details of pH measurement and the preparation of standard acid and alkali solutions the reader is referred to a companion volume A Albert and E P Serjeant s *The Determination of Ionization Constants* 1971 Although the aims of the book are essentially practical it also deals in some detail with those theoretical aspects considered most helpful to an understanding of buffer applications We have cast our net widely to include pH buffers for particular purposes and for measurements in non aqueous and mixed solvent systems In recent years there has been a significant expansion in the range of available buffers particularly for biological studies largely in consequence of the development of many zwitterionic buffers by Good et al 1966 These are described in Chapter 3 **Determination of PH** Roger Gordon Bates,1973 *Isoelectric Point Determination* Merrell Beyer,1936 **Measurement of PH** Richard E. Lawn,Florence Elizabeth Prichard,2003 Whether pH is being used to test a sample against a legal requirement or specification as part of an analytical method for monitoring and controlling a reaction as a process control in the chemical industry or for the environmental monitoring of waste and effluents it is important that all pH measurements are carried out in a logical and consistent manner paying careful attention to experimental procedures in order to obtain reliable results This guide provides scientists with the knowledge of how to do just that first by outlining

the principles of pH measurement and buffer solutions pH meters and electrodes are then discussed including selection criteria and the care of electrodes Finally sections on making pH measurements and uncertainty are followed by a set of practical exercises Measurement of pH is one of the Practical Laboratory Skills Training Guides a series that aims to make achieving best practice easy These invaluable manuals will enable both experienced and inexperienced staff to get the essential basics of any experiment right simply by following the clear and easy to use instructions provided The guides are written by experienced scientists and include minimal theory plenty of practical exercises in order to assess competence and trouble shooting information Other titles are Measurement of Mass Measurement of Volume High Performance Liquid Chromatography and Gas Chromatography *Flinn Scientific Explain and Predict--Practice Free Response Questions for AP Chemistry* Flinn Scientific,2014 **Standard Buffer Solutions** George G. Manov,1947 For the purpose of calibrating pH meters such as the hydrogen calomel and the glass calomel type it is necessary to have on hand a number of buffer solutions of certified pH value covering the ranges of temperature and pH over which it is desired to work Of the large number of compounds both organic and inorganic known to the chemist comparatively few however have the requisite properties such as uniformity and reproducibility of composition freedom from deliquescence or efflorescence ease of preparation and purification and good buffer capacity which would make them acceptable as buffer materials In the past thirty years the concept of the term pH underwent a change from that of the simple definition of $\text{pH} = \log_{10} \frac{1}{[\text{H}^+]}$ to the more modern definition of $\text{pH} = \log_{10} a_{\text{H}^+}$ when it was realized that the potential of the hydrogen electrode was a measure of the activity of the hydrogen ion rather than of its concentration Paradoxically however the activity of single ions cannot be determined from experiment without the use of some simplifying assumptions it is therefore impossible to establish a pH scale on a rigorous thermodynamic basis This point will be discussed later The problem of purification of salts is one primarily of analytical chemistry and the materials can be tested to determine their relative purity and reproducibility by any self consistent method

pH and Buffer Theory-A New Approach H. Rilbe,1996-11-21 pH and buffer theory is a central concept in synthetic and industrial processes Since speed and type of reaction are usually governed by the pH of the reaction medium an understanding of the theory is important in predicting results This book addresses all the key aspects of pH and buffer theory and makes the equations and corresponding graphs easy to grasp *Colorimetric Determination of PH Without Buffer Solutions* P. H. Symons,1934

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