

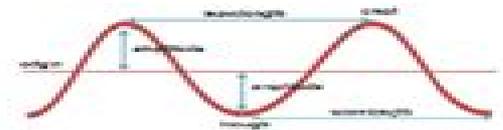
1. Light and Quantized Energy: The Wave Nature of Light

The Weaknesses of Rutherford's Model

1. Did not explain how the atom's electrons are arranged in the space around the nucleus
2. Did not address the question of why the negatively charged electrons are not pulled into the atom's positively charged particles
3. Did not account for the differences and similarities in chemical behavior among the various elements

Wavelength

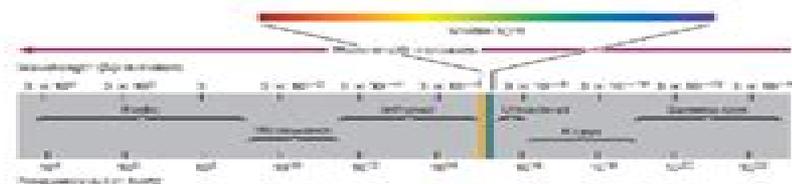
- Wavelength: the distance between two waves
- Frequency: the number of waves that pass a given point per second
- Wavelength and frequency do not have an effect on the amplitude of a wave
- You can measure wavelengths: anywhere with the same name points
 - Crest to crest
 - Trough to trough
 - Origin to Origin
- Wavelength is usually measured in meters, cm, and nm



Relationships -

- All waves travel at the same speed $\rightarrow 3.00 \times 10^8$ m/s
- The longer the wavelength, the lower the frequency (inverse relationship)
 - When frequency is high, the wavelength is short
 - When the wavelength is long, the frequency is low
- The more energy, the higher the frequency
- Electromagnetic Wave Relationship: $c = \lambda \nu$
 - $c \rightarrow$ the speed of light in a vacuum = 3.00×10^8 m/s
 - $\lambda \rightarrow$ wavelength (lambda)
 - $\nu \rightarrow$ frequency (nu)

Electromagnetic Radiation



Light And Quantized Energy Study Guide Answer

R Bogdan



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