

Energy and Chemical Change

Section 15.1 Energy

pages 516–522

Practice Problems

pages 519–521

1. A fruit and oatmeal bar contains 142 nutritional Calories. Convert this energy to calories.

$$142 \text{ Calories} = 142 \text{ kcal}$$

$$142 \text{ kcal} \times \frac{1000 \text{ cal}}{1 \text{ kcal}} = 142,000 \text{ cal}$$

2. An exothermic reaction releases 86.5 kJ. How many kilocalories of energy are released?

$$86.5 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = 20.7 \text{ kcal}$$

3. **Challenge** Define a new energy unit, named after yourself, with a magnitude of one-tenth of a calorie. What conversion factors relate this new unit to joules? To Calories?

$$\text{Unit X} = 0.1 \text{ cal}$$

$$1 \text{ cal} = 4.184 \text{ J}$$

$$\text{X} = (0.1 \text{ cal})(4.184 \text{ J/cal}) = 0.4184 \text{ J}$$

$$1 \text{ cal} = 0.001 \text{ Calorie}$$

$$\text{X} = (0.1 \text{ cal})(1 \text{ Cal}/1000 \text{ cal}) = 0.0001 \text{ Calorie}$$

4. If the temperature of 34.4 g of ethanol increases from 25.0°C to 78.8°C, how much heat has been absorbed by the ethanol? Refer to Table 15.2.

$$q = c \times m \times \Delta T$$

$$q = 2.44 \text{ J/(g}\cdot\text{°C)} \times 34.4 \text{ g} \times 53.8\text{°C} = 4.52 \times 10^3 \text{ J}$$

5. A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance? Identify the substance among those listed in Table 15.2 on page 520.

$$q = c \times m \times \Delta T$$

$$c = \frac{q}{m\Delta T} = \frac{(5696 \text{ J})}{(155 \text{ g})(40.0 - 25.0\text{°C})} = 2.45 \text{ J/(g}\cdot\text{°C)}$$

The specific heat is very close to the value for ethanol.

6. **Challenge** A 4.50-g nugget of pure gold absorbed 276 J of heat. The initial temperature was 25.0°C. What was the final temperature?

$$q = c \times m \times \Delta T$$

$$\Delta T = \frac{q}{cm} = \frac{(276 \text{ J})}{(0.129 \text{ J/g}\cdot\text{°C})(4.50 \text{ g})} = 475\text{°C}$$

$$\Delta T = T_f - T_i = T_f - 25.0\text{°C} = 475\text{°C}$$

$$T_f = 5.00 \times 10^2\text{°C}$$

Section 15.1 Assessment

page 522

7. **Explain** how energy changes from one form to another in an exothermic reaction. In an endothermic reaction.

Chemical potential energy changes to heat in exothermic reactions and the heat is released. In endothermic reactions, heat is absorbed and changed to chemical potential energy.

8. **Distinguish** between kinetic and potential energy in the following examples: two separated magnets; an avalanche of snow; books on library shelves; a mountain stream; a stock-car race; separation of charge in a battery.

Two separated magnets illustrate potential energy. In a snow avalanche, positional potential energy is changing to kinetic energy. Books on a shelf illustrate positional potential energy. As water races down a mountain stream, positional potential energy is changing to kinetic energy. In a stock-car race, chemical potential energy is being changed to kinetic energy. The separation of charge in a battery illustrates electrical potential energy.

9. **Explain** how the light and heat of a burning candle are related to chemical potential energy.

Chemical potential energy, contained in the candle, is changed to energy in the form of light and heat and released as the chemical combustion reaction takes place.

Matter And Change Chapter 8 2 Study Guide Key

**Katherine J. Denniston, Denniston
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Table of Contents Matter And Change Chapter 8 2 Study Guide Key

1. Understanding the eBook Matter And Change Chapter 8 2 Study Guide Key
 - The Rise of Digital Reading Matter And Change Chapter 8 2 Study Guide Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Matter And Change Chapter 8 2 Study Guide Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Matter And Change Chapter 8 2 Study Guide Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Matter And Change Chapter 8 2 Study Guide Key
 - Personalized Recommendations

- Matter And Change Chapter 8 2 Study Guide Key User Reviews and Ratings
- Matter And Change Chapter 8 2 Study Guide Key and Bestseller Lists
- 5. Accessing Matter And Change Chapter 8 2 Study Guide Key Free and Paid eBooks
 - Matter And Change Chapter 8 2 Study Guide Key Public Domain eBooks
 - Matter And Change Chapter 8 2 Study Guide Key eBook Subscription Services
 - Matter And Change Chapter 8 2 Study Guide Key Budget-Friendly Options
- 6. Navigating Matter And Change Chapter 8 2 Study Guide Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Matter And Change Chapter 8 2 Study Guide Key Compatibility with Devices
 - Matter And Change Chapter 8 2 Study Guide Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Matter And Change Chapter 8 2 Study Guide Key
 - Highlighting and Note-Taking Matter And Change Chapter 8 2 Study Guide Key
 - Interactive Elements Matter And Change Chapter 8 2 Study Guide Key
- 8. Staying Engaged with Matter And Change Chapter 8 2 Study Guide Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Matter And Change Chapter 8 2 Study Guide Key
- 9. Balancing eBooks and Physical Books Matter And Change Chapter 8 2 Study Guide Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Matter And Change Chapter 8 2 Study Guide Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Matter And Change Chapter 8 2 Study Guide Key
 - Setting Reading Goals Matter And Change Chapter 8 2 Study Guide Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Matter And Change Chapter 8 2 Study Guide Key
 - Fact-Checking eBook Content of Matter And Change Chapter 8 2 Study Guide Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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