

Name: _____ Date _____ Period: _____

Molarity & Molality Notes and Practice

Answer the questions below. **SHOW ALL WORK**, including units!! Watch your significant digits and **CIRCLE YOUR ANSWERS**.

Molarity

Just a reminder, molarity is one of the many ways to measure concentration or the strength of a solution. When using molarity to measure concentration you must follow the formula below and then put a capital M at the end of your answer to let the world know you used the molarity formula.

$$M = \frac{\text{moles of solute}}{\text{Liters of solvent}}$$

1. Calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution.
2. What is the molarity of a solution containing 325 g of NaCl dissolved in 750. mL of solution?
3. 140 g of KCl is dissolved in 600. mL of water. What is the molarity?
4. 724.4 g of ammonium phosphate in 4500 mL of alcohol. What is the molarity of the solution?
5. You are making 2.2 L of 3.1 M silver nitrate solution. How many moles of solute are there?
6. How many grams of MgCl_2 are needed to make 700 mL of a 1.4 M solution?
7. 93.2 g of copper (II) sulfate is mixed into 290. mL of water. What is the molarity?

Molality

Molality is an additional way to measure the strength or concentration of a solution. It is abbreviated with a little m and is calculated only slightly differently than molarity. Here is the formula.

$$m = \frac{\text{moles of solute}}{\text{kg of solvent}}$$

Molarity And Molality Practice Worksheet Answers

**David W. Oxtoby, H. Pat Gillis, Laurie J.
Butler**

Molarity And Molality Practice Worksheet Answers:

Problems and Solutions to Accompany Chang and Lovett's Physical Chemistry for the Biosciences, Second Edition Mark D. Marshall, Helen O. Leung, 2025-12-16 Containing detailed solutions to each of the even numbered problems from the second edition of Raymond Chang and Charles M Lovett s Physical Chemistry for the Biosciences this manual gives students a strong pedagogical framework for self study Nothing can better help students understand difficult concepts than working through and solving problems Featuring detailed solutions to each of the even numbered problems from the second edition of Raymond Chang and Charles M Lovett s Physical Chemistry for the Biosciences this manual gives students fresh insights into concepts that may elude them in the lecture hall and a strong pedagogical framework for self study The authors approach each solution with the conversational style used in their classrooms designed to teach students problem solving techniques rather than simply handing out answers The book features illustrative figures and diagrams throughout

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