

Molarity:

- a quantitative description of solution concentration.
- Abbreviated M

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

Problems: Show all work and circle your final answer.

1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required?

$$4.00 \text{ M} = \frac{\text{moles of solute}}{12.0 \text{ L}} \quad \text{moles of solute} = \boxed{48.0 \text{ mol}}$$

2. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M?

$$? \text{ L} = 250 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.25 \text{ L}$$

$$0.150 \text{ M} = \frac{\text{moles of solute}}{0.25 \text{ L}} \quad \text{moles of solute} = \boxed{0.038 \text{ mol}}$$

3. What is the molarity of a solution of HNO₃ that contains 12.6 grams HNO₃ in 1.0 L of solution?

$$? \text{ mol HNO}_3 = 12.6 \text{ g HNO}_3 \times \frac{1 \text{ mol HNO}_3}{63.0 \text{ g HNO}_3} = 0.200 \text{ mol HNO}_3$$

$$M = \frac{0.200 \text{ mol HNO}_3}{1.0 \text{ L}} = \boxed{0.200 \text{ M}}$$

4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?

$$0.700 \text{ M} = \frac{\text{moles of solute}}{0.250 \text{ L}} \quad \text{moles of solute} = 0.175 \text{ mol}$$

$$? \text{ g KNO}_3 = 0.175 \text{ mol KNO}_3 \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = \boxed{17.7 \text{ g KNO}_3}$$

Molarity Calculations Worksheet Answers And Work

Nick Morris



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web mar 1 2022 the following is intended as a guide for beginners in the airbrush technique it is about the equipment of the airbrush the different painting techniques and how you create your first and future images the explanations should begin with the introduction of the airbrush technique what is airbrush