

THE MOLE: LAB ACTIVITY

Station 1:

A) Determine the mass of one drop of water by measuring the mass of 50 drops of water.

| | |
|---|-----------|
| Mass of dish (or beaker) | = _____ g |
| Mass of dish (or beaker) + 50 drops water | = _____ g |
| Mass of 50 drops water | = _____ g |
| Mass of 1 drop of water | = _____ g |

B) How many molecules are there in one drop of water?

Remember 1 mole of a substance is 6.02×10^{23} particles (atoms or molecules).
Therefore, 6.02×10^{23} molecules of water will weigh 18.02 g/mol

1. Find # of moles of water for 1 drop $n = (\text{mass})/(\text{Molar Mass})$
2. # of molecules in one drop of water $= n_{\text{water}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}}$
 $= ? \text{ molecules}$

Station 2

Calculate the number of copper atoms in a penny.

1. Find # of moles of Cu $n = (\text{mass})/(\text{Molar Mass})$
2. # of Cu atoms in a penny $= n_{\text{Cu}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$
 $= ? \text{ atoms}$

Station 3

A. Calculate the number of moles in one spoon of sugar. Assume that sugar is made up entirely of sucrose molecules ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$).

$$M_{\text{sucrose}} = 342.34 \text{ g/mol}$$

$$n_{\text{sugar}} = \frac{M_{\text{sucrose}}}{M_{\text{sucrose}}}$$
$$= \underline{\hspace{2cm}} \text{ mol}$$

B. Calculate the number of molecules of sugar in one spoon of sugar.

$$\begin{aligned} \# \text{ of molecules sucrose} &= n_{\text{sucrose}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= \underline{\hspace{2cm}} \text{ mol} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} \\ &= ? \text{ molecules} \end{aligned}$$

Mole Lab Student Guide Answers

Kathleen Armour



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Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **General,**

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