

The Mole Test 3

Instructions: There have to identical units in pairs, then make sure all the moles are balanced in ratios. Molar mass can be used to convert grams to moles or moles to grams. Molar volume can be used to convert moles to volume or volume to moles. Always use the correct units (grams, moles, volume). Always use the correct units (grams, moles, volume). Always use the correct units (grams, moles, volume).

Part 1: Balance 10 Questions

1) How many particles are in 1.00 mole of Ag₂SO₄?

1.00 mol	1.00 × 10 ²³ particles	1.00 × 10 ²³ particles Ag ₂ SO ₄
	1. mol Ag ₂ SO ₄	

2) How many moles are 1.00 × 10²³ particles of CO₂?

1.00 × 10 ²³ particles	1. mol CO ₂	1.00 × 10 ²³ particles
	1.00 × 10 ²³ particles	1.00 × 10 ²³ particles

3) How many particles are in 1.00 mole of H₂O?

1.00 mol	1.00 × 10 ²³ particles	1.00 × 10 ²³ particles H ₂ O
	1. mol	

4) How many moles are 1.00 × 10²³ particles of H₂O?

1.00 × 10 ²³ particles	1. mol H ₂ O	1.00 × 10 ²³ particles
	1.00 × 10 ²³ particles	1.00 × 10 ²³ particles

Part 2: Balance 10 Questions

Calculate the molar masses for the following chemical formulas. Show your work!

<u>Chemical Formula</u>	<u>Molar Mass (g/mol)</u>
1) H ₂ O	H ₂ = 2 × 1.01 g/mol O = 16.00 g/mol H ₂ O = 18.02 g/mol

Mole Test Answer

Jacob Samuel Orleans, Glenn A. Sealy

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