

# The Mole Study Guide

The Mole

**Molar Mass**

1. Calculate the molar mass of the following compounds:

- (a)  $\text{H}_2\text{O}$
- (b)  $\text{CO}_2$
- (c)  $\text{NaCl}$
- (d)  $\text{CH}_4$
- (e)  $\text{H}_2\text{SO}_4$
- (f)  $\text{Ca(OH)}_2$
- (g)  $\text{Al}_2\text{O}_3$
- (h)  $\text{Fe}_2\text{O}_3$
- (i)  $\text{KNO}_3$
- (j)  $\text{HNO}_3$

**Molar Volume**

2. Calculate the volume of the following gases at STP:

- (a) 2 moles of  $\text{O}_2$
- (b) 0.5 moles of  $\text{H}_2$
- (c) 1 mole of  $\text{CO}_2$
- (d) 3 moles of  $\text{N}_2$
- (e) 1.5 moles of  $\text{CH}_4$

**Molar Concentration**

3. Calculate the molar concentration of the following solutions:

- (a) 2 moles of  $\text{NaCl}$  in 1 liter of solution
- (b) 0.5 moles of  $\text{H}_2\text{SO}_4$  in 2 liters of solution
- (c) 1 mole of  $\text{KNO}_3$  in 0.5 liters of solution
- (d) 3 moles of  $\text{Fe}_2\text{O}_3$  in 1.5 liters of solution

4. Calculate the number of moles of the following substances:

- (a) 36 g of  $\text{H}_2\text{O}$
- (b) 44 g of  $\text{CO}_2$
- (c) 58.5 g of  $\text{NaCl}$
- (d) 16 g of  $\text{CH}_4$
- (e) 98 g of  $\text{H}_2\text{SO}_4$
- (f) 74 g of  $\text{Ca(OH)}_2$
- (g) 102 g of  $\text{Al}_2\text{O}_3$
- (h) 160 g of  $\text{Fe}_2\text{O}_3$
- (i) 101 g of  $\text{KNO}_3$
- (j) 63 g of  $\text{HNO}_3$

5. Calculate the mass of the following substances:

- (a) 2 moles of  $\text{O}_2$
- (b) 0.5 moles of  $\text{H}_2$
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# Mole Unit Study Guide Answers

**Lauren Gardner**



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