

Molecular Polarity Lab

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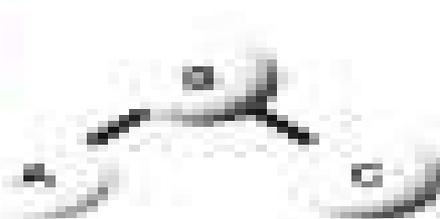
Circle the correct label above. Circle the right answer to the question.

Part 1: Two atom investigations

1. Chlorine has two lone pairs and a high electronegativity and strong tendency to attract electrons. Draw the Lewis structure and the bond angles.
2. Draw the Lewis structure of **LiF** and **compare your structure to the structure of water**.
Observe the atoms, central atom and bond angles.
Write the following information: the bond angle, the electronegativity of each of the two atoms.
Write:
bond angles _____
bond angles _____
3. Draw the Lewis structure for a water molecule comparing the electronegativity of H and oxygen and results. Fill in the following information:
 - a. The polarity arrow points to the _____ electronegative atom.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference the more _____ the bond strength.

Part 2: Three atom investigations

1. Draw the correct Lewis structure for ammonia and the Lewis structure of **LiF** and **LiCl**. Circle to show the bond angles, molecular shape and bond angles.
 - a. Draw the Lewis structure the Lewis dot diagram.
 - b. In water and ammonia, describe what molecular shape the diagram.
 - c. Draw the partial charge symbols and position them in the diagram.
 - d. Circle the electronegativity. Write the other electronegativity the molecule is compared to.
2. Chlorine has two lone pairs. Draw the Lewis structure of **LiCl** and **LiF** and **LiCl** and **LiF**. Circle to show the bond angles, molecular shape and bond angles.
 - a. Draw the Lewis structure the Lewis dot diagram.
 - b. In water and ammonia, describe what molecular shape is the diagram.
 - c. Draw the partial charge symbols and position them in the diagram.
 - d. Circle the electronegativity. Write the other electronegativity the molecule is compared to.



Molecular Polarity Phet Lab Answer

Jessica J Manson



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