

Molecular Polarity and Electronegativity: A PHET Online Inquiry lab

Name _____ Section _____ Date _____

PHET Simulation: Molecular Polarity

Directions:

- Go to the following website: <https://phet.colorado.edu/en/simulation/molecular-polarity>
- Click the word **SIMULATIONS** prominently at the screen and choose **Molecular Polarity** from the drop down menu.
- Read the simulation's **Molecular Polarity** and click it to open it.
- Click the play button.

Or go directly to this address in your browser:
<https://phet.colorado.edu/en/simulation/molecular-polarity>

First, click on the top space that says, "Three atoms"

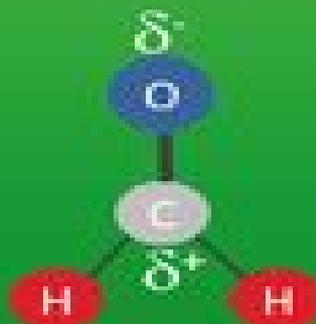
Experiment with the simulator. Try out the controls, click and drag things to move them around, click buttons and try to get an idea of how the simulator works. When you finish, push the reset button  and then answer the questions below.

1) Which atom has the highest electronegativity?

Click the check boxes for bond dipole and partial charges 

2) Why do you think that the molecule dipole arrow points upward? Consider the bond dipoles.

3) Is this molecule polar? Explain. Consider the partial charges of each atom.



Name _____ Section _____ Date _____

Molecular Polarity Questions

1) How does the molecular dipole arrow in the simulator tell you if molecule is polar or nonpolar and indicates the partial positive and negative charges on in the molecule?

2) How do the bond dipole arrows relate to whether or not a molecule is polar?

3) When you turn on the dipoles that follow questions 1 and 2, why do the molecules interact with the dipoles. Think in the way that it did?

4) When you turn on the dipoles that follow question 3, why do the molecules interact with the dipoles. Think in the way that it did?

A CHEMISTRY INQUIRY LAB

WORKED ANSWER KEY INCLUDED

GRADING GUIDE INCLUDED

Molecule Polarity Phet Lab Teachers Answer Key

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