

Atoms and Molecules

Reading Comprehension Passages

Name: _____ Date: _____

Multiple Choice Questions

Read each statement and choose the correct answer.

1. Oxygen gas is a
A. Compound
B. Element
C. Mixture
D. Solution
2. Atoms in solids
A. Move freely
B. Are far apart
C. Vibrate in place
D. Do not move
3. Molecules are
A. Larger than atoms
B. Smaller than atoms
C. Smaller
D. Similar
4. Why proposed atomic theory?
A. Matter
B. Matter
C. Matter
D. Matter
5. Periodic table is arranged by
A. Mass
B. Size
C. Atomic number
D. None

Name: _____ Date: _____

Fill in the Blanks

1. Atoms are the _____ of matter.
2. Electrons have a _____ charge.
3. A molecule is made of _____ atoms.
4. The nucleus is at the _____ of the atom.
5. Chemical bonds hold _____ together.
6. Matter has _____.
7. Atomic number equals number of _____.
8. Molecules in gases have _____.
9. John Dalton proposed _____ theory.
10. Atoms have _____ being neutral.

Name: _____ Date: _____

True or False

Read each statement and circle if it is True or False.

1. Atoms are the smallest particles of matter.
2. Molecules are made of atoms.
3. Protons have a positive charge.
4. All atoms are the same.
5. Chemical reactions rearrange atoms.
6. Matter is continuous.
7. Electrons move around the nucleus.
8. Molecules exist in solids only.
9. Atoms freely combine matter.
10. Atoms are important in matter.

Name: _____ Date: _____

Reading Passages #7

Chemical Bonds

Chemical bonds are forces that hold atoms together in molecules. Atoms bond to become more stable. There are different types of bonds, such as covalent and ionic bonds. In covalent bonding, atoms share electrons. In ionic bonding, one atom transfers electrons to another. Bonds determine the shape and properties of molecules. Strong bonds create stable compounds, while weak bonds may break easily. Chemical bonding explains how atoms form molecules and why substances behave differently. Without chemical bonds, molecules would not exist, and life as we know it would be impossible.

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