

# Chemistry: Molality

Directions: Using these formulas, solve the following problems.

$$\text{molality} = \text{mol solute} / \text{kg solvent}$$

$$1 \text{ amu} = 1 \text{ gram} / \text{mol}$$

$$\Delta T_{\text{boiling}} = (i \cdot \text{number of ions}) (\text{molality}) (k_b)$$

$$\Delta T_{\text{freezing}} = (\text{number of ions}) (\text{molality}) (k_f)$$

$k_b$  and  $k_f$  are constants for each individual solvent.

Key



Problem 1: What is the molality of a solution containing 1.23 g of  $\text{Hg}(\text{CN})_2$  dissolved in 1.37 g of methanol? Answer: 3.55 m

Problem 2: What is the molality of a solution containing 16.8 g of NaCl dissolved in 356 g of water? Answer: 0.606 m

Problem 3: How many grams of NaOH would be required to produce a 3.57 m solution with 473 g of water? Answer: 67.6 g

Problem 4: How many grams of KCl would you need to dissolve in 114 g of vinegar in order to prepare a 1.96 m solution? Answer: 16.6 g

Problem 5: Calculate the molality of 5.97 g NaCl dissolved in 471 g of water.  
Answer: 0.217 m

Problem 6: Calculate the molality of 8.91 g  $\text{Na}_2\text{CO}_3$  dissolved in 595 g of water.  
Answer: 0.341 m

Problem 7: Calculate the molality of 8.44 g KBr dissolved in 112 g of water.  
Answer: 0.633 m

Problem 8: What are the new boiling and freezing points for a solution with 581 g KBr dissolved in 845 grams of water? ( $k_b$  water =  $0.512^\circ \text{C/m}$  and  $k_f$  water =  $1.86^\circ \text{C/m}$ )  
Answer: new freezing point =  $-13.4^\circ \text{C}$       new boiling point =  $106^\circ \text{C}$

Problem 9: What are the new boiling and freezing points for a solution with 391 g NaOH dissolved in 1029 grams of water? ( $k_b$  water =  $0.512^\circ \text{C/m}$  and  $k_f$  water =  $1.86^\circ \text{C/m}$ )  
Answer: new freezing point =  $-36.3^\circ \text{C}$       new boiling point =  $107^\circ \text{C}$

# Molality Answer Key

**Marian DeWane, Heather Hattori**



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**of Chemistry** David E. Goldberg, 2001 Iterations, II Russell Batt, John W. Moore, 1987 **Chemistry** James C. Hill, 2003

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