

Molarity

Molarity is just a unit of concentration that indicates how much solute we have in a solution. It is defined as shown below.

$$M = \frac{\text{mols of solute}}{\text{L. of solution}}$$

Thus, we need to remember the definitions of solute, solvent and solution:

solute: the substance dissolved in a solvent to form a solution; it is normally the component of a solution present in the smaller amount

solvent: the dissolving medium of a solution; it is normally the component of a solution present in the greater amount

solution: a mixture of substances that has a uniform composition; a homogeneous mixture

So, let's get some practice with this new unit!

Convert the following Molarity expressions into a conversion factor. The first one is done as an example.

1. $6.45 \text{ M NaCl} = \frac{6.45 \text{ mols NaCl}}{1 \text{ L. of NaCl sol'n}}$

4. $0.0234 \text{ M NaHCO}_3 = \underline{\hspace{2cm}}$

2. $10.2 \text{ M HCl} = \underline{\hspace{2cm}}$

5. $1.58 \times 10^{-9} \text{ M Na}_2\text{SO}_4 = \underline{\hspace{2cm}}$

3. $3.42 \text{ M NaOH} = \underline{\hspace{2cm}}$

6. $1.58 \times 10^{-4} \text{ M KCH}_3\text{CO}_2 = \underline{\hspace{2cm}}$

Now, let's try to do some calculations with Molarity! Remember to always write out M as the mol/L so you can see all the units.

7. How many moles of NaCl are in 2.50 liters of 6.45 M NaCl?

$$\frac{6.45 \text{ mols NaCl}}{1 \text{ L. NaCl}} \times 2.50 \text{ L. NaCl} = 16.1 \text{ mols NaCl}$$

8. How many moles of HCl are in 0.54 liters of 10.2 M HCl?

9. How many liters of NaOH are in 1.3 moles of 3.42 M NaOH?

Molality M Answer Key

Ralph H. Petrucci, William S. Harwood

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