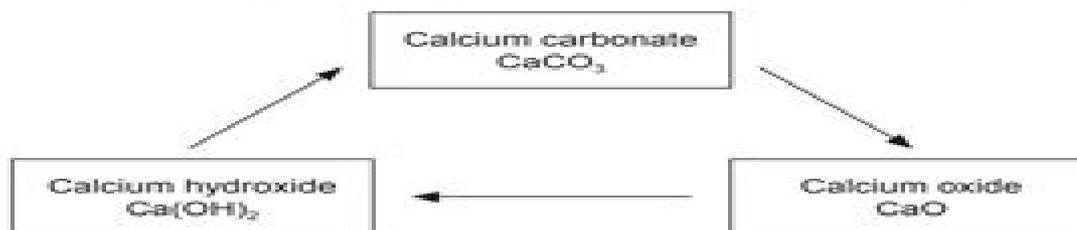


Answer all questions. Answers must be written within the answer boxes provided.

1. Limestone can be converted into a variety of useful commercial products through the lime cycle. Limestone contains high percentages of calcium carbonate, CaCO_3 .

$\frac{17}{20}$



- (a) Calcium carbonate is heated to produce calcium oxide, CaO .



Calculate the volume of carbon dioxide produced at STP when 555 g of calcium carbonate decomposes. Use sections 2 and 6 of the data booklet.

[2] 7

$\text{mol of CaCO}_3: \frac{555}{100} = 5.55$ $\text{mol of CO}_2 = 5.55$
 ratio: 1:1
 $\text{volume} = 5.55 \times 22.7 = 126.0$ dm^3
 calc mistake $\approx 126.0 \text{ dm}^3$

(This question continues on the following page)



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log linear analysis is a technique used in statistics to examine the relationship between more than two categorical variables the technique is used for both hypothesis testing and model building in both these uses models are tested to find the most parsimonious i e least complex model that best accounts for the variance in the observed frequencies a pearson s chi square test could be used instead of log linear analysis but that technique only allows for two of the variables to be c