



Name:

Date:

Student Exploration: Calorimetry Lab

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- The Latin word *calor* means "heat," and *meter* comes from the Greek word meaning "to measure." What do you think a **calorimeter** does?

measuring the heat of chemical reactions or physical changes as well as heat capacity

- Where have you heard the word **calorie** before? What do you think a calorie is?

calorie, a unit of energy or heat variously defined. the calorie was originally defined as the amount of heat required at a pressure of 1 standard atmosphere to raise the temperature of 1 gram of water 1 degree celsius.

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



- On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

- What was the **Final temperature** of the copper and the water?
- How much did the temperature of the copper change?
- How much did the temperature of the water change?

34.96

down 55.04

down 165.04

- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Explore Learning Calorimetry Lab Gizmo Answer Key

Judd E. Hollander



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