

Quiz & Worksheet - Colligative Properties and Raoult's Law

1. What happens to vapor pressure when you add a solute to a solution?

- It lowers the vapor pressure.
- It has no effect.
- It raises the vapor pressure.
- It causes the reaction to reach equilibrium.

2. What is osmotic pressure?

- It is the minimum pressure that must be applied to a solution to stop osmosis from happening.
- It is the maximum pressure that must be applied to a solution to stop osmosis from happening.
- It is the maximum pressure of the vapor over a liquid at equilibrium.
- It is the minimum pressure of the vapor over a liquid at equilibrium.

3.

9.0g of urea (a nonelectrolyte) is dissolved in 360 g water to create a solution. The temperature of the solution is 25.0°C. If the vapor pressure of water at 25.0°C is 23.80 mmHg, calculate the vapor pressure of the urea solution at 25.0°C in mmHg.

The molar mass of urea 60g/mol and water is 18g/mol.

- 0.59
- 23.91
- 23.61
- 3.57

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Explore Learning Colligative Properties Answers

Albert A Gayle



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In a global taken by screens and the ceaseless chatter of instantaneous conversation, the melodic elegance and mental symphony created by the prepared word frequently diminish into the backdrop, eclipsed by the constant sound and disruptions that permeate our lives. Nevertheless, set within the pages of **Explore Learning Colligative Properties Answers** a marvelous fictional treasure full of fresh thoughts, lies an immersive symphony waiting to be embraced. Crafted by an outstanding musician of language, this captivating masterpiece conducts visitors on a mental journey, well unraveling the hidden songs and profound affect resonating within each carefully crafted phrase. Within the depths with this touching review, we can examine the book is central harmonies, analyze its enthralling writing design, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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