

Student Exploration: Nuclear Decay (ANSWER KEY)

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Student Exploration: Nuclear Decay

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

The chart below gives the locations, charges, and approximate masses of three **subatomic particles**. The approximate mass of each particle is given in universal mass units (*u*).

| Particle | Location | Charge | Approximate mass |
|----------|----------|--------|------------------|
| Proton | Nucleus | 1+ | 1 u |
| Neutron | Nucleus | 0 | 1 u |
| Electron | Orbitals | 1- | 0 u |

1. The **mass number** of an atom is equal to the sum of protons and neutrons in the nucleus. A helium atom has 2 protons and 2 neutrons. What is the mass number of this atom?

1. The **atomic number** of an element is equal to the number of protons in each atom of the element. All helium atoms have 2 protons. What is the atomic number of helium?

Gizmo Warm-up

While most atoms are stable, some are **radioactive**, which means that they have a tendency to undergo spontaneous **nuclear decay**. The decay of radioactive atoms generally results in the **emission of particles and/or energy**.

Explore Learning Nuclear Decay Answer Key

Richard Bailey



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