

## Student Exploration: Titration

<b>Activity A:</b> <b>Acids and bases</b>	<u>Get the Gizmo ready:</u> <ul style="list-style-type: none"> <li>• Click <b>Reset</b>. Select <b>1.0 M HNO<sub>3</sub></b> for the <b>Burette</b> and <b>Mystery NaOH</b> for the <b>Flask</b>.</li> <li>• Select <b>Phenolphthalein</b> for the <b>Indicator</b>.</li> <li>• You will need a scientific calculator for this activity.</li> </ul>	
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**Question: How do acids and bases interact in solution?**

1. **Describe:** The equation for the reaction of nitric acid (HNO<sub>3</sub>) and sodium hydroxide (NaOH) is shown on the bottom right of the Gizmo.
  - A. What are the reactants in this reaction? HNO<sub>3</sub>, NaOH
  - B. What are the products of this reaction? NaNO<sub>3</sub>, H<sub>2</sub>O
2. **Measure:** A **titration** can be used to determine the concentration of an acid or base by measuring the amount of a solution with a known concentration, called the **titrant**, which reacts completely with a solution of **unknown concentration**, called the **analyte**. The point at which this occurs is called the **equivalence point**.

Carefully add HNO<sub>3</sub> into the flask until the phenolphthalein begins to lose its color. Stop adding HNO<sub>3</sub> when the color change is permanent.

- A. How much (HNO<sub>3</sub>) was required to cause the indicator to change color? 8.9mL
  - B. What can you say about the pH before and after the last drop of HNO<sub>3</sub> was added?  
 The pH was above 8.2 when it was pink before the last drop, after the last drop it is less than 8.2
3. **Explore:** Click **Reset** and change the indicator to **Bromthymol blue**. Add exactly 8.8 mL of HNO<sub>3</sub> to the flask.
    - A. What does the color of the indicator tell you about the current pH of the flask?  
 Since it is blue, it tells me that the pH is above 7.6
    - B. Add one more drop of HNO<sub>3</sub>. What does the color tell you about the pH now?  
 Since it is yellow, I know that the pH is below 6.0

<b>Activity B:</b>	<u>Get the Gizmo ready:</u>	
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